

## Flame Test Lab

### Activity Key

#### Background

The heat from a laboratory burner will cause the ions of some elements to give off light. Electrons will absorb the heat energy from the flame and will “jump” to a higher energy level. When the electrons return to their original energy levels, this absorbed energy is released as light. Different elements absorb and release different amounts of energy and thus produce different kinds of light. This light may be visible or invisible, infrared or ultraviolet radiation. Analysis of this light can be used to identify elements and molecules. In today’s lab, we will observe the visible range of light emission.

#### Objectives

1. Perform a flame test to identify the characteristic color of metal ions.
2. Identify the metal ion in an unknown solution.
3. Calculate the energy of emitted photons.

#### Materials: (per lab group)

10 ml dilute solutions of the following:

Barium Chloride  
Calcium Chloride  
Lithium Chloride  
Potassium Chloride  
Sodium Chloride  
Strontium Chloride  
Unknown Solution

7 Wood Splints  
1 Bunsen Burner  
Safety Goggles & Aprons

#### Note:

If chloride compounds are not available, metal nitrate compounds may be substituted.

Use dilute or approximately 0.1 M solutions

#### Unknowns:

Number the beakers 1-12.  
Add solutions as follows:

- |                       |                        |
|-----------------------|------------------------|
| 1. Strontium Chloride | 7. Calcium Chloride    |
| 2. Calcium Chloride   | 8. Strontium Chloride  |
| 3. Potassium Chloride | 9. Barium Chloride     |
| 4. Barium Chloride    | 10. Potassium Chloride |
| 5. Lithium Chloride   | 11. Sodium Chloride    |
| 6. Sodium Chloride    | 12. Lithium Chloride   |

#### Procedure

1. Put on your safety goggles and an apron.
2. Obtain a wood splint and the barium chloride solution.
3. Dip the wood splint into the solution and allow it to soak for 1-2 minutes.
4. Light Bunsen Burner. USE CAUTION!
5. Remove the splint from the solution and hold the end of the splint into the flame. Do not burn the splint.
6. Observe and record the color of the flame.
7. Repeat steps 2-6 using the remaining chloride solutions.
8. Obtain an unknown solution from your teacher. Record the number of your unknown.
9. Soak a wood splint in the unknown solution.
10. Hold the splint in the flame and record the color of the flame that is produced.
11. Using your data, identify the metal ion in your unknown solution.

# Chemistry Flame Test Lab Answers

**Mary Pride**



## Chemistry Flame Test Lab Answers:

*Chemistry in the Laboratory* James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations *Laboratory Manual for Principles of General Chemistry* Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures **Lab Experiments for Modern Chemistry** Tzimopoulo, 1990 *Holt Chemistry* Ralph Thomas Myers, 2004 **Laboratory Exercises to Accompany First Principles of Chemistry** Raymond Bedell Brownlee, 1908 *General Chemistry* Ralph H. Petrucci, 2007 *General Chemistry Principles and Modern Applications* is recognized for its superior problems lucid writing precision of argument and precise and detailed treatment of the subject Popular and innovative features include Feature Problems follow up A and B Practice Exercises to accompany every in chapter Example Focus On application boxes and new Keep in Mind marginal notes Every new copy of the Ninth Edition comes with a Student MediaPak which includes access to the Companion Website with GradeTracker available at <http://www.prenhall.com/petrucci> the Student Accelerator CD and the Virtual ChemLab Workbook and CD This package includes Basic Media Pack Wrap Companion WEBSITE Grade Tracker Access Code Card Virtual ChemLab General Chemistry Student Lab Manual Workbook **Big Book of Home Learning** Mary Pride, 1991-07 Learn at home with exciting products for all school subjects New *Chemistry* Karen Timberlake, 1999 Suitable for one or two term lab courses covering general organic and biological chemistry this new edition written by Karen Timberlake features many improvements to the insightful experiments that have made it the leading lab manual Each experiment encourages critical thinking with laboratory goals discussion of related concepts clear instructions new pre lab questions and comprehensive report pages Forty one experiments illustrate the basic principles of chemistry *Experiments in General Chemistry* Kenneth W. Whitten, Carl B. Bishop, Kenneth D. Gailey, 1992 *General, Organic, and Biological Chemistry Study Guide and Selected Solutions* Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems *A Laboratory Text Book of Practical Chemistry* William George Valentin, 1871 **Exercises in Chemistry, Systematically Arranged to Accompany "An Elementary Study of Chemistry", Introductory College Course, by William Pherson and William Edwards Henderson** William McPherson, 1925 *The Journal of Industrial and Engineering Chemistry*, 1913 **Chemistry** John S. Phillips, Cheryl

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Gradwohl Laboratory Digest ,1964-07     **Instructors Manual to Lab Manual** Ralph Petrucci,William  
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web 3 if one zero of the quadratic polynomial  $x^2 + 3x + k$  is 2 then the value of  $k$  is a 10 b 10 c 5 d 5 1 4 graphically the pair of equations given by  $6x + 3y = 10$  and  $2x + y = 9$  represents two lines which are a intersecting at exactly one point b parallel c coincident d intersecting at exactly two points 1

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web if a point  $(x, y)$  is equidistant from the points  $(9, 8)$  and  $(17, 8)$  then i  $x = y$  ii  $x = 13$  iii  $y = 13$  iv  $x = y = 13$  18 case study based 2 scale factor and similarity scale factor a scale drawing of an object is the same shape as the object but a different size the scale of a drawing is a comparison of the length used on a drawing to