

# "CHEMICAL BONDING"



8 electron



6 electron

↓  
Short octet  
Group IIIA element  
in compound form  
show short octet



10 electron  
↓  
extended octet

- a)  $\text{NH}_3$
- b)  $\text{SO}_2$
- c)  $\text{AlCl}_3$
- d) B and C

- SHORT OCTETS ARE ALWAYS LEWIS ACID.
- EXTENDED OCTETS MAY OR MAY NOT BE LEWIS BASE.

## ENERGETICS OF BOND FORMATION:

### Formation

$$\Delta H = -ve$$

Energy released

So overall energy less

So stability occurs.

Bond length

or  
Bond distance

or  
compromise bond distance.

### Breakage

$$\Delta H = +ve$$

Heat absorbed.

Overall energy increases

Stability loses.

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**JA Banks**



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