

ACIDS, BASES and Salts

Important Terms

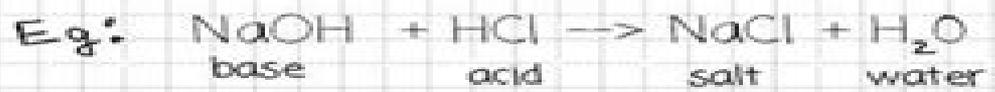
Acids - Substances that dissociate in water to give H^+ or H_3O^+ ions are called acids [according to Arrhenius concept].

Bases - Substances that dissociate in water to give OH^- (hydroxy) ions are called bases [according to Arrhenius concept].

Neutralization reaction:



Salt + Water



water-soluble bases = alkalis

pH represents the strength of an acid/base in terms of Hydrogen ion concentration

Strong acid + Strong Base = Neutral Salt + Water

Strong Acid + Weak Base = Acidic Salt + Water

Weak Acid + Strong Base = Basic Salt + Water

Conjugate base - ion/radical left after either H^+ ion (proton) is released after acid dissociates
Conjugate acid - ion/radical left after base accepts proton (H^+ ion)

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