

## CHAPTER 2 WORKSHEET: CHEMISTRY OF LIFE

### 1. Chapter 2, Section 1: Nature of Matter

- a. Matter: anything that occupies space. Ex: Desk, Human, Air, ect.
- b. Atom: Smallest part of matter. It **can not** be broken down into smaller pieces. Composed of 3 particles: Protons (+), Electrons (-), and Neutrons (n)
- c. Element: Substance that is only made of one kind of atom. Ex: Hydrogen, Oxygen, ect.
  - i. C, H, N, O, P, S
- d. Compound: Two or more different atoms or elements joined together. Ex: Salt =
- e. Covalent Bond: Two or more atoms sharing electrons to form a molecule. Covalent bonds make molecules, ex: Water, Carbon Dioxide, and Oxygen Gas
- f. Hydrogen Bond: A weak chemical attraction between **polar** molecules. Ex: Two water molecules
- g. Ionic Bond: Gain or Loss of electrons that create a molecule
- h. Ion: An atom that has gained (-) electrons or (loss) electrons (+)

### 2. Chapter 2, Section 2: Water and Solutions

- a. How does water help maintain homeostasis?
  - i. Heats slowly
  - ii. Retains heat
  - iii. Carries Heat away through water vapor (sweating)
  - iv. Universal solvent (anything except for a polar substance can dissolve in it)
- b. What is the difference between adhesion & cohesion?
  - i. Adhesion: different substances
  - ii. Cohesion: alike substances
- c. What is the pH range for an Acid? Base?
  - i. Acid = 0 – 6.9
  - ii. Neutral = 7
  - iii. Base = 7.1 – 14
- d. What type of substance is an example of an acid? Of a base?
  - i. Acids: Lemon, Vinegar, Milk
  - ii. Bases: Household cleaning material

### 3. Chapter 2, Section 3: Chemistry of Cells

- a. Carbohydrates: Organic substances, made up of carbon, hydrogen and oxygen atoms. Always in the ratio of 1:2:1. Key source of energy.
  - i. They are made up of Monosaccharide: Single sugars (glucose)
  - ii. Disaccharides are double sugars: Sucrose
  - iii. Polysaccharides: Starch and cellulose

# Chapter Test Chemistry Of Life Answer Key

**Gary A. Thibodeau, Kevin T. Patton**



## Chapter Test Chemistry Of Life Answer Key:

**Chapter Resource 2 Chemistry of Life Biology** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004

*Chemistry and Life* John William Hill, Dorothy M. Feigl, Stuart J. Baum, 1993      **Chemistry and Life** John W. Hill, 1997

**Life** William K. Purves, 2001 Authoritative thorough and engaging Life The Science of Biology achieves an optimal balance of scholarship and teachability never losing sight of either the science or the student The first introductory text to present biological concepts through the research that revealed them Life covers the full range of topics with an integrated experimental focus that flows naturally from the narrative This approach helps to bring the drama of classic and cutting edge research to the classroom but always in the context of reinforcing core ideas and the innovative scientific thinking behind them Students will experience biology not just as a litany of facts or a highlight reel of experiments but as a rich coherent discipline

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