

## Prentice Hall Physical Science

### Chapter 4 Atomic Structure

#### 4.1 Studying Atoms

- A. Ancient Greek Models of Atoms
- Democritus believed all matter was made of extremely small particles that could not be divided
  - he called the particles atoms (from the Greek atomos meaning indivisible)
  - he thought there were different types of atoms with different properties (liquids had round, smooth atoms and solids were rough and prickly)
- B. Dalton's Atomic Theory
- he proved atoms exist by discovering that compounds have a fixed composition
  - Dalton's Theory:
    - all elements are composed of atoms
    - all atoms of the same element have the same mass, and atoms of different elements have different masses
    - compounds contain atoms of more than one element
    - in a particular compound, atoms of different element always combine in the same way
  - he pictured atoms as solid spheres
  - in time it was found that not all of this was completely correct
- C. Thomson's Model of the Atom
- some objects have a positive or negative charge and objects with the same charge repel and objects with opposite charges attract
  - Thomson used an electric current to learn more about atoms
  - he passed an electric current through a gas and a glowing beam appeared
  - the beam was repelled by a negatively charged metal plate (like charges repel and opposite charges attract)
  - Thomson concluded that the beam must be made of negatively charged particles, smaller than atoms.
  - since the atom was known to be electrically neutral, he developed a model in which negative charges were evenly scattered throughout an atom filled with a positively charged pudding like material
  - called the "Plum Pudding" model (or chocolate chip cookie dough)
- D. Rutherford's Atomic Theory
- Rutherford discovered positively charged, fast moving particles called alpha particles
  - he shot some of these particles a piece of gold foil, thinking they would go straight through
  - when he actually performed the experiment, he found that the positively charged alpha particles were greatly deflected by the gold (figure 7, p. 104)
  - he determined that the alpha particles must have come very close to another positively charged particle causing them to be repelled and put off course
  - this led him to believe that the positive charge in an atom is not evenly distributed, but concentrated in a very small central area he called the nucleus
  - his model of the atom is a small, positively charged nucleus surrounded by negatively charged particles

# Chapter Assessment Prentice Hall Physical Science

**Thomas Griffiths**



**Chapter Assessment Prentice Hall Physical Science:**

## Adopting the Beat of Term: An Psychological Symphony within **Chapter Assessment Prentice Hall Physical Science**

In a world consumed by monitors and the ceaseless chatter of instantaneous interaction, the melodic splendor and emotional symphony produced by the prepared word usually diminish in to the backdrop, eclipsed by the constant noise and interruptions that permeate our lives. However, nestled within the pages of **Chapter Assessment Prentice Hall Physical Science** a charming literary treasure filled with organic feelings, lies an immersive symphony waiting to be embraced. Constructed by a wonderful musician of language, this fascinating masterpiece conducts readers on a mental trip, well unraveling the concealed tunes and profound affect resonating within each cautiously constructed phrase. Within the depths of this moving review, we will investigate the book is central harmonies, analyze their enthralling publishing fashion, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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