

## Section 12.1

### OBJECTIVE 1

Write complete balanced redox reactions that are indicated.

1. \_\_\_\_\_
  - a. \_\_\_\_\_
    - (a) balance in terms of all the atoms and products.
    - (b) indicate oxidation-reduction numbers of reactants and products.
    - (c) identify all atoms of each element in each compound as oxidized or reduced.
    - (d) identify all electrons transferred, including a net flow.
  - b. \_\_\_\_\_
    - (a) identify the oxidizing agent and the reducing agent.
    - (b) the mass of product produced when a certain mass of reactant decomposes.
    - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
    - (d) the balanced chemical equation based on the net ionic method.
2. \_\_\_\_\_
  - a. \_\_\_\_\_
    - (a) identify all atoms and oxidation numbers of all reactants and products.
    - (b) energy released in the reaction.
    - (c) identify all electrons of all elements in the reaction.
    - (d) number of moles of hydrogen that will react with a certain quantity of oxygen.
  - b. \_\_\_\_\_
    - (a) identify the oxidizing agent and the reducing agent.
    - (b) number of moles of hydrogen that will react with a certain quantity of oxygen.
    - (c) balanced chemical equation.
    - (d) balanced equation.

3. Write the net ionic equation for the following redox reaction:



a. Balance the equation:



b. Write the net ionic equation of  $\text{Mg} + \text{Fe}^{2+}$ .

c. How many moles of  $\text{Fe}^{2+}$  ions are reduced by 1.0 mole of  $\text{Mg}$ ?

d. How many moles of  $\text{Fe}^{2+}$  ions are reduced by 1.0 mole of  $\text{Mg}$ ?

e. How many grams of  $\text{Fe}^{2+}$  ions are reduced by 1.0 mole of  $\text{Mg}$ ?

# Chapter 9 Review Stoichiometry Answers Section 9 1 Bing

**LL Leslie**



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