

CHAPTER 8 REVIEW

Chemical Equations and Reactions

SHORT ANSWER Answer the following questions in the space provided.

2. _____ According to the law of conservation of mass, the total mass of the reacting substances is
- (a) always more than the total mass of the products.
 - (b) always less than the total mass of the products.
 - (c) sometimes more and sometimes less than the total mass of the products.
 - (d) always equal to the total mass of the products.**

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 2. In the reaction described by the word equation
- $$\text{sodium} + \text{water} \rightarrow \text{sodium hydroxide} + \text{hydrogen gas}$$

- a. sodium hydroxide is a product.
- b. sodium is a product.
- c. water is a reactant.
- d. Both (a) and (c).**

- _____ 3. Which of the following indicates that a chemical equation is balanced?

- a. The numbers of atoms of each element are the same on both sides of the equation.**
- b. All of the coefficients are the same.
- c. The numbers of molecules on each side of the equation are equal.
- d. The sums of the coefficients on each side of the equation are equal.

- _____ 4. From a complete and correctly written chemical equation, you can obtain the

- a. chemical formulas of the reactants and products.
- b. relative amounts of the reactants and products.
- c. physical states of the reactants and products.
- d. All of the above.**

- _____ 5. To balance a chemical equation, you must

- a. adjust the subscripts.
- b. adjust the number of products.
- c. adjust the number of reactants.
- d. adjust the coefficients.**

- _____ 6. The formation of a solid product in a chemical reaction is represented by the symbol

- a. (aq).
- b. (g).
- c. (l).
- d. (s).**

Chapter 8 Test Chemical Equations And Reactions

Modern Chemistry

Xiang Xie



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