

Temperature

°C, K

measure of average KE
of motion of particles

Heat

kJ, kcal (Cal) [1 kcal = 4.184 kJ]

measure of total energy transferred from
an object of high E → low E.

Note: A change in T is accompanied by a transfer of heat energy.

Specific heat (c or c_p):

(specific for each substance)

amount of energy required to raise the temperature of 1 g of that substance by 1°C.

change in heat = specific heat x mass x change in T

$$q = c_p \times m \times \Delta T$$

Units: heat absorbed or released, in J;

ΔT in °C or K; m in g; c in J/g°C

Molar heat capacity = heat capacity in J/mol°C instead of J/g°C

Don't forget: the size of 1°C = 1 K, so K can be used in this calculation in place of °C.

Examples

- How much heat energy is released to your body when a cup of hot tea containing 200. g of water is cooled from 65.0°C to body temperature, 37.0°C? (c_p of H_2O = 4.18 J/g°C)
- A 4.50-g nugget of pure gold absorbed 276 J of heat. What was the final temperature of the gold if the initial temperature was 25.0°C? (c_p of gold = 0.129 J/g°C)
- A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance?

Chapter 6 Thermochemistry Test

Ralph H. Petrucci, Robert K. Wismer



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Table of Contents Chapter 6 Thermochemistry Test

1. Understanding the eBook Chapter 6 Thermochemistry Test
 - The Rise of Digital Reading Chapter 6 Thermochemistry Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Thermochemistry Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 6 Thermochemistry Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Thermochemistry Test
 - Personalized Recommendations
 - Chapter 6 Thermochemistry Test User Reviews and Ratings
 - Chapter 6 Thermochemistry Test and Bestseller Lists
5. Accessing Chapter 6 Thermochemistry Test Free and Paid eBooks
 - Chapter 6 Thermochemistry Test Public Domain eBooks
 - Chapter 6 Thermochemistry Test eBook Subscription Services
 - Chapter 6 Thermochemistry Test Budget-Friendly Options

6. Navigating Chapter 6 Thermochemistry Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 6 Thermochemistry Test Compatibility with Devices
 - Chapter 6 Thermochemistry Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 6 Thermochemistry Test
 - Highlighting and Note-Taking Chapter 6 Thermochemistry Test
 - Interactive Elements Chapter 6 Thermochemistry Test
8. Staying Engaged with Chapter 6 Thermochemistry Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 6 Thermochemistry Test
9. Balancing eBooks and Physical Books Chapter 6 Thermochemistry Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 6 Thermochemistry Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 6 Thermochemistry Test
 - Setting Reading Goals Chapter 6 Thermochemistry Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 6 Thermochemistry Test
 - Fact-Checking eBook Content of Chapter 6 Thermochemistry Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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