

AP Chemistry Chemical Kinetics Worksheet Answers

1.	Half lives	Reactant Fraction Remaining
	1	$\frac{1}{2}$
	2	$\frac{1}{4}$
	3	$\frac{1}{8}$
	4	$\frac{1}{16}$
	5	$\frac{1}{32}$

2. 2nd Order; 1st Order

3. a. True b. True

c. False; As [Reactant] decreases the rate will decrease d. True

4. a. $-1/2 \Delta[\text{O}_3]/\Delta t = 1/3 \Delta[\text{O}_2]/\Delta t$
b. $-1/2 \Delta[\text{HOF}]/\Delta t = 1/2 \Delta[\text{HF}]/\Delta t = \Delta[\text{O}_2]/\Delta t$
c. $-\Delta[\text{N}_2]/\Delta t = -1/3 \Delta[\text{H}_2]/\Delta t = 1/2 \Delta[\text{NH}_3]/\Delta t$

5. a. rate = $k[\text{NO}_2]^2[\text{CO}]^0$ or $k[\text{NO}_2]^2$
b. rate = 1/4 the original value

6. The initial rate $R = k[\text{B}]_0^2$

7. a. $6.0 \times 10^{-3} / 1.2 \times 10^{-2} = (0.10/0.20)^X$
 $X = 1$ so it is 1st order with respect to A
 $1.1 \times 10^{-1} / 1.2 \times 10^{-2} = (0.15/0.05)^X$
 $X = 2$ so it is 2nd order with respect to B
Rate Law = $k[\text{A}][\text{B}]^2$

b. Overall it is 3rd order

c. Probably not a single step since a 3-body collision would have to occur.

8. a. half life = $\ln(1/2) / -5.0 \times 10^{-4} \text{ s}^{-1} = 1.4 \times 10^3 \text{ s}$
 $-\ln(0.5) / 5.0 \times 10^{-4} \text{ s}^{-1} = 1.4 \times 10^3 \text{ s}$
b. Assign an original value of 1.00M
 $\ln(0.1/1.0) = -5.0 \times 10^{-4} \text{ s}^{-1} t$
 $t = -\ln(0.1) / 5.0 \times 10^{-4} \text{ s}^{-1} = 4.6 \times 10^3 \text{ s}$

9. After 30.min (one half-life)

$$P_{\text{HOF}} = \frac{1}{2} (100. \text{mmHg}) = 50. \text{mmHg}$$

$$P_{\text{HF}} = 50. \text{mmHg}$$

$$P_{\text{O}_2} = 25. \text{mmHg}$$

Calculating the rate constant k ,

$$\ln([\text{HOF}]_t / [\text{HOF}]_0) = -kt$$

$$\ln(0.5) = -k(30. \text{min})$$

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Zhe Xu



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Carolina Science and Math Carolina Biological Supply Company, 2003

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CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

Rates and Mechanisms of Chemical Reactions W. C. Gardiner (Jr.), William Cecil Gardiner, 1969

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