

Unit 1: Chemistry of Life

Elements

- **Elements:** are substances that cannot be broken down into simpler substances by chemical means.
- Oxygen (O), carbon (C), hydrogen (H), and nitrogen (N).
- These four elements are used to build biological molecules, such as carbohydrates, proteins, lipids, and nucleic acids. They are also used to form storage compounds and cells in all organisms.
- Some elements are known as **trace elements** because they are required by an organism only in very small quantities. They include iron (Fe), iodine (I), and copper (Cu).
- Other elements are present but in smaller quantities.

Subatomic Particles

- **Atoms** are the unit of life and are the building blocks of the physical world.
 - Protons are positively charged (+) particles
 - Neutrons are uncharged particles.
 - Electrons are negatively charged (-) particles
- Some atoms have the same number of protons but differ in the number of neutrons in the nucleus. These are called isotopes.

Compounds

- Consists of two or more elements
- The atoms of a compound are held together by chemical bonds, which may be ionic bonds, covalent bonds, or hydrogen bonds.
- An **ionic bond** is formed between two atoms when one or more electrons are transferred from one atom to the other.
- The charged forms of the atoms are called ions.
- A **covalent bond** is formed when electrons are shared between atoms. If the electrons are shared equally between the atoms, the bond is called non-polar covalent. If the electrons are shared unequally, the bond is called polar covalent.

Water: The Versatile Molecule

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