

Inorganic Chemistry

Lesson 1: Inorganic Chemistry

Inorganic Chemistry - deals with **MATTER** its composition, properties and its reactions.

Matter - is anything that occupies space and has mass.

- All the objects in the universe, since they occupy space and have mass, are matter.
- The mass of an object is the quantity of matter that it contains.

Classification of Matter:

1. **PURE SUBSTANCES / SUBSTANCES** have fixed compositions and distinct properties.
 - **ELEMENTS** - are substances that cannot be decomposed into simpler substances by chemical means. They are the basic substances out of which all matter is composed.
 - Metal
 - Nonmetal
 - Metalloid

COMPOUNDS - are the substances composed of two or more elements united chemically in definite proportions by mass. They can be decomposed by chemical means into simpler elements.

- Acid
 - Base
 - Salt
2. **MIXTURES** are composed of two or more kinds of matter where each substance retains its own identity and properties thus it can be separated by physical means. Compositions of mixtures can vary continuously.
- **HOMOGENEOUS MIXTURES (1 phase / SOLUTIONS)** - have uniform composition throughout. Solutions are classified under this group, i.e. *vinegar, brine, syrup.*

→ **HETEROGENEOUS MIXTURES (2/ SUSPENSIONS)** - are substances whose composition, properties, and appearance differ throughout, i.e. sand in water, chocolate chip ice cream, fruit salad.

- ◆ **silverware** - a homogeneous mixture (alloy)

Properties of Matter - refer to the characteristics that enable one to distinguish one substance to another.

1. **Physical properties:** color, odor, taste, hardness, physical state, melting point, boiling point, electrical conductivity, malleability, density, solubility, specific gravity.
2. **Chemical properties:** rusting metal, burning metal, burning fuel, milk turning sour

Physical States Matter

1. **Solid** - the molecules that make up a solid are arranged in regular, repeating patterns. They are held firmly in place but can vibrate within a limited area.
2. **Liquid** - the molecules that make up a liquid flow easily around one another. They are kept from flying apart by attractive forces between them. Liquids assume the shape of their containers.
3. **Gas** - the molecules that make up a gas fly in all directions at great speeds. They are so far apart that attractive forces between them are insignificant.
4. **Plasma** - at the very high temperatures of stars, atoms lose their electrons. The mixture of electrons and nuclei that results is the plasma state of matter.

Changes of Matter

1. **Physical change** does not involve a change of one of matter into another; no change in the fixed composition of the substance.

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