

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

\_\_\_\_\_ a. ionic bond  
 \_\_\_\_\_ b. covalent bond  
 \_\_\_\_\_ c. metallic bond

2. Use the electronegativity values shown in Figure 5-20, on page 151 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ a. H—F	_____ d. H—H
_____ b. Na—Cl	_____ e. H—C
_____ c. H—O	_____ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

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4.  $\text{H}_2\text{S}$  and  $\text{H}_2\text{O}$  have similar structures and their central atoms belong to the same group. Yet  $\text{H}_2\text{S}$  is a gas at room temperature and  $\text{H}_2\text{O}$  is a liquid. Use bonding principles to explain why this is true.

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# Holt Modern Chemistry Chapter 6 Review Packet Answers

**J Elliott**



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