

Ideal gas law - problems and solutions

1. Ideal gases in a closed container initially have volume V and temperature T . The final temperature is $5/4T$ and the final pressure is $2P$. What is the final volume of gases.

Known:

Initial volume (V_1) = V

Initial temperature (T_1) = T

Final temperature (T_2) = $5/4 T$

Initial pressure (P_1) = P

Final pressure (P_2) = $2P$

Wanted: Final volume (V_2)

Solution:

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{P V}{T} = \frac{(2P) V_2}{5/4 T}$$

$$\frac{V}{1} = \frac{2 V_2}{5/4}$$

$$5/4 V = 2 V_2$$

$$V_2 = \frac{5V}{4} \cdot \frac{2}{1}$$

$$V_2 = \frac{5V}{4} \times \frac{1}{2}$$

$$V_2 = \frac{5V}{8}$$

2. Determine the volume of 2.00 moles of gases (ideal gas) at STP. STP = Standard Temperature and Pressure.

Known:

Moles of gas (n) = 2 moles

Standard temperature (T) = $0^\circ\text{C} = 0 + 273 = 273$ Kelvin

Standard pressure (P) = 1 atm = 1.013×10^5 Pa

Universal gas constant (R) = 8.315 Joule/mole.Kelvin

Wanted: Volume of gases (V)

Solution:

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