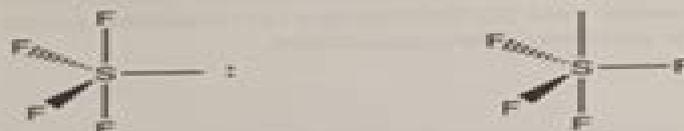


Post Laboratory Assignment

Answer the following on a separate sheet of paper.

- Using the data collected for PF_5 , SF_6 and ClF_3 , what structural comments can be made about molecules with more than four groups (charge clouds) around the central atom?
- Select which structure is preferred and discuss why?



- To assess the stability of the axial vs. equatorial positions for a trigonal bipyramidal structure, a computer program was used to determine the energy for SF_4 with the lone pair in either position (axial or equatorial). The data obtained is shown below:

Equatorial: Energy = -795.14 au

Axial: Energy = -795.088 au

- Using the conversion factor $627.5096 \text{ kcal} = 1 \text{ au}$, what is the energy difference in kcal between the two structures?
 - According to these data, which structure is more stable (more negative)?
 - Does this agree with your explanation from the previous question? Why or why not?
- Below is a table that lists dipole moments and bond angles calculated by a computer program for three molecules. Discuss why the bond angles differ and describe the significance for the dipole moment difference.

Molecule	Dipole moment	Bond Angle
Methane	0.00 Debye	109.47 °
Ammonia	1.84 Debye	107.57°
Water	2.15 Debye	105.93 °

- Below is a table that lists dipole moments for three molecules. Discuss the ionic and covalent character of each compound using the dipole moment values.

Molecule	Dipole moment
H_2	0.00 Debye
HF	1.94 Debye
LiH	5.95 Debye

Experiment 11 Molecular Models Answers

AW Rasmussen



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