

# CHEMISTRY

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## PERIODIC TABLE OF THE ELEMENTS

The periodic table shows elements grouped into color-coded categories:
 

- Hydrogen (H):** Atomic number 1, atomic weight 1.008. Diagram shows a single 1s orbital.
- s-subshell:** Diagram shows a single spherical orbital.
- p-subshell:** Diagram shows three dumbbell-shaped orbitals.
- d-subshell:** Diagram shows five complex, multi-lobed orbitals.

### ATOMIC STRUCTURE

Atomic Number,  $Z$ , of an element is the number of protons in the nucleus. For neutral atoms,  $Z = +$  the number of electrons. The mass number,  $A$ , is the sum of the number of protons and the number of neutrons. Isotopes of an element have the same  $Z$ , different  $A$ . The atomic mass is the weighted average of isotopes. The atomic weight is the weighted average of isotopes for a given element.

### ATOMIC QUANTUM NUMBERS AND ORBITALS

Orbitals and quantum numbers are used to describe all atoms, molecules, ions, and ions. A quantum number describes the energy, size, and orbital properties. The main quantum number,  $n$ , is the principal quantum number. The orbital quantum number,  $l$ , is the azimuthal quantum number. The magnetic quantum number,  $m_l$ , is the magnetic quantum number. The spin quantum number,  $m_s$ , is the spin quantum number.

Orbital	$n$	$l$	$m_l$	$m_s$
1s	1	0	0	$\pm \frac{1}{2}$
2s	2	0	0	$\pm \frac{1}{2}$
2p	2	1	0, $\pm 1$	$\pm \frac{1}{2}$
3s	3	0	0	$\pm \frac{1}{2}$
3p	3	1	0, $\pm 1$	$\pm \frac{1}{2}$
3d	3	2	0, $\pm 1, \pm 2$	$\pm \frac{1}{2}$

### NUCLEAR CHEMISTRY

Nuclear reactions occur when particles (neutrons, protons, alpha particles, beta particles, gamma rays, etc.) interact with nuclei. The energy released in nuclear reactions is much greater than in chemical reactions. The mass number,  $A$ , is the sum of the number of protons and neutrons. The atomic number,  $Z$ , is the number of protons. The nuclear equation is balanced for both  $A$  and  $Z$ .

Types of processes:
 

- Radioactive decay:**
  - $\alpha$  decay:  $n \rightarrow p + \alpha$
  - $\beta^-$  decay:  $n \rightarrow p + e^- + \bar{\nu}_e$
  - $\beta^+$  decay:  $p \rightarrow n + e^+ + \nu_e$
  - $\gamma$  decay:  $\text{excited nucleus} \rightarrow \text{ground state} + \gamma$
- Transmutation:**
  - Artificial:  $p + \text{target} \rightarrow \text{product} + \text{particle}$
  - Natural:  $\alpha + \text{target} \rightarrow \text{product} + \text{particle}$

### MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe all atoms, molecules, ions, and ions. A quantum number describes the energy, size, and orbital properties. The main quantum number,  $n$ , is the principal quantum number. The orbital quantum number,  $l$ , is the azimuthal quantum number. The magnetic quantum number,  $m_l$ , is the magnetic quantum number. The spin quantum number,  $m_s$ , is the spin quantum number.

Orbital	$n$	$l$	$m_l$	$m_s$
1s	1	0	0	$\pm \frac{1}{2}$
2s	2	0	0	$\pm \frac{1}{2}$
2p	2	1	0, $\pm 1$	$\pm \frac{1}{2}$
3s	3	0	0	$\pm \frac{1}{2}$
3p	3	1	0, $\pm 1$	$\pm \frac{1}{2}$
3d	3	2	0, $\pm 1, \pm 2$	$\pm \frac{1}{2}$

Atomic configuration:  $1s^2 2s^2 2p^6 3s^2 3p^4$

Atomic configuration:  $1s^2 2s^2 2p^6 3s^2 3p^4 3d^5$

Detailed periodic table showing element symbols, atomic numbers, and names. The table is color-coded by groups and periods.

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