

## Chapter 6

# Equilibrium

### Question 6.47:

It has been found that the pH of a 0.01M solution of an organic acid is 4.15. Calculate the concentration of the anion, the ionization constant of the acid and its  $pK_a$ .

### Answer 6.47:

Let the organic acid be HA.



Concentration of HA = 0.01 M pH = 4.15

$$-\log[H^+] = 4.15$$

$$[H^+] = 7.08 \times 10^{-5}$$

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

Now,

$$[H^+] = [A^-] = 7.08 \times 10^{-5}$$

$$[HA] = 0.01$$

Then,

$$K_a = \frac{(7.08 \times 10^{-5})(7.08 \times 10^{-5})}{0.01} = 5.01 \times 10^{-7}$$

$$pK_a = -\log K_a = -\log(5.01 \times 10^{-7}) = 6.3001$$

### Question 6.48:

Assuming complete dissociation, calculate the pH of the following solutions:

(i) 0.003 M HCl

(ii) 0.005 M NaOH

(iii) 0.002 M HBr

(iv) 0.002 M KOH

### Answer 6.48:

(i) 0.003M HCl:



Since HCl is completely ionized,

$$[H_3O^+] = [HCl] = 0.003$$

Now,

$$pH = \log[H_3O^+] = -\log(.003) = 2.52$$

Hence, the pH of the solution is 2.52.

# Chemistry Chapter 6 Solutions Manual

**Robert F. Francis**



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