

# CHEMISTRY



**Properties of Matter**

- Matter is anything that has mass and occupies space.
- Matter is made up of particles.
- Particles are very small.
- Particles are continuously moving.
- Particles attract each other.

**States of Matter**

- Solid: Particles are closely packed and have a fixed shape and volume.
- Liquid: Particles are less closely packed and have a fixed volume but no fixed shape.
- Gas: Particles are far apart and have no fixed shape or volume.

**Changes of State**

- Melting: Solid to liquid.
- Freezing: Liquid to solid.
- Boiling: Liquid to gas.
- Condensation: Gas to liquid.
- Sublimation: Solid to gas.
- Deposition: Gas to solid.

**Atomic Structure**

Atoms are the smallest particles that cannot be created or destroyed in a chemical reaction. They are made up of three sub-particles: electrons, protons, and neutrons.

**Electrons**

- Discovered by Cathode ray experiment.
- Charge:  $-1.6 \times 10^{-19}$  C.
- Mass:  $9.1 \times 10^{-31}$  kg.

**Protons**

- Discovered by Rutherford's alpha particle scattering experiment.
- Charge:  $+1.6 \times 10^{-19}$  C.
- Mass:  $1.67 \times 10^{-27}$  kg.

**Neutrons**

- Discovered by Chadwick.
- Charge: 0.
- Mass:  $1.67 \times 10^{-27}$  kg.

**Atomic Structure (Continued)**

Atoms are electrically neutral. The number of protons is equal to the number of electrons. The nucleus is made up of protons and neutrons. Electrons are present in shells around the nucleus.

**Atomic Number**

The atomic number of an element is the number of protons in the nucleus of an atom of that element. It is denoted by the symbol  $Z$ .

Element	Atomic Number (Z)	Number of Protons	Number of Electrons
Hydrogen	1	1	1
Helium	2	2	2
Lithium	3	3	3
Boron	5	5	5
Carbon	6	6	6
Nitrogen	7	7	7
Oxygen	8	8	8
Fluorine	9	9	9
Neon	10	10	10

**Atomic Structure (Continued)**

The atomic number of an element is equal to the number of protons in the nucleus. It is also equal to the number of electrons in a neutral atom.



**Atomic Structure (Continued)**

Electrons are arranged in shells around the nucleus. The shells are numbered 1, 2, 3, 4, 5, 6, 7 from innermost to outermost. The maximum number of electrons in a shell is given by  $2n^2$ , where  $n$  is the shell number.

**Atomic Structure (Continued)**

The outermost shell is called the valence shell. The number of electrons in the valence shell is called the valency of the element.



**Atomic Structure (Continued)**

The atomic number of an element is equal to the number of protons in the nucleus. It is also equal to the number of electrons in a neutral atom.

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**Chung Rodgers**



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## **Table of Contents Quick Study Laminated Reference Guides**

1. Understanding the eBook Quick Study Laminated Reference Guides
  - The Rise of Digital Reading Quick Study Laminated Reference Guides
  - Advantages of eBooks Over Traditional Books
2. Identifying Quick Study Laminated Reference Guides
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a Quick Study Laminated Reference Guides
  - User-Friendly Interface
4. Exploring eBook Recommendations from Quick Study Laminated Reference Guides
  - Personalized Recommendations
  - Quick Study Laminated Reference Guides User Reviews and Ratings
  - Quick Study Laminated Reference Guides and Bestseller Lists
5. Accessing Quick Study Laminated Reference Guides Free and Paid eBooks
  - Quick Study Laminated Reference Guides Public Domain eBooks
  - Quick Study Laminated Reference Guides eBook Subscription Services
  - Quick Study Laminated Reference Guides Budget-Friendly Options
6. Navigating Quick Study Laminated Reference Guides eBook Formats

- ePub, PDF, MOBI, and More
- Quick Study Laminated Reference Guides Compatibility with Devices
- Quick Study Laminated Reference Guides Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Quick Study Laminated Reference Guides
  - Highlighting and Note-Taking Quick Study Laminated Reference Guides
  - Interactive Elements Quick Study Laminated Reference Guides
- 8. Staying Engaged with Quick Study Laminated Reference Guides
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Quick Study Laminated Reference Guides
- 9. Balancing eBooks and Physical Books Quick Study Laminated Reference Guides
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Quick Study Laminated Reference Guides
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Quick Study Laminated Reference Guides
  - Setting Reading Goals Quick Study Laminated Reference Guides
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Quick Study Laminated Reference Guides
  - Fact-Checking eBook Content of Quick Study Laminated Reference Guides
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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