

## Practicing with Isotopes

1. Complete the following questions. Assume all the atoms are neutral.



Element: \_\_\_\_\_

# of protons: \_\_\_\_\_

# of neutrons: \_\_\_\_\_



Element: \_\_\_\_\_

# of protons: \_\_\_\_\_

# of neutrons: \_\_\_\_\_



Element: \_\_\_\_\_

# of protons: \_\_\_\_\_

# of neutrons: \_\_\_\_\_



Element: \_\_\_\_\_

# of protons: \_\_\_\_\_

# of neutrons: \_\_\_\_\_



Element: \_\_\_\_\_

# of protons: \_\_\_\_\_

# of neutrons: \_\_\_\_\_



Element: \_\_\_\_\_

# of protons: \_\_\_\_\_

# of neutrons: \_\_\_\_\_



Element: \_\_\_\_\_

# of protons: \_\_\_\_\_

# of neutrons: \_\_\_\_\_



Element: \_\_\_\_\_

# of protons: \_\_\_\_\_

# of neutrons: \_\_\_\_\_



Element: \_\_\_\_\_

# of protons: \_\_\_\_\_

# of neutrons: \_\_\_\_\_

2. Calculate the average atomic mass for the following element from the given information about the relevant isotopes:

Isotope 1: Mass = 35 amu

Isotope 2: Mass = 37 amu

Abundance = 75.53%

Abundance = 24.47%

Which element is this? \_\_\_\_\_

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Pg 91

**TD Snyder**



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