

CHAPTER

02

SOLUTION

- Solution** : A homogeneous mixture of two or more substances is known as solution.
- Solute** : The substance present in smaller amount in a solution is called solute.
- Solvent** : The substance present in larger amount in a solution is called solvent.
- Binary solution** : These solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
- Ternary solutions** : Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
- Aqueous solution** : When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
- Non-aqueous solution** : When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
- Weight Fraction = $\frac{\text{weight of solute}}{\text{weight of solution}}$
- Weight percentage

$$= \frac{\text{weight of solute}}{\text{weight of solution}} \times 100$$

Weight of solute present in 100 grams of solution is called weight percentage

Volume of solute present in 100 ml of the solution is called volume percentage
- Volume percentage

$$= \frac{\text{Volume of solute}}{\text{Volume of solution}} \times 100$$
- Volume Fraction = $\frac{\text{Volume of solute}}{\text{Volume of solution}}$
- Molarity (M)

$$= \frac{\text{No of moles of solute}}{\text{Volume of solution in liters}}$$

(or) $M = \frac{W}{GMW} \times \frac{1}{V(lit)}$
- Molarity = $\frac{\text{No of moles of solute}}{\text{Volume of solution in ml}} \times 1000$
- Molarity = $\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{V \text{ in ml}}$
 (or) $M = \frac{W}{GMW} \times \frac{1000}{V(ml)}$
- Weight of solute = $\frac{M \times GMW \times V \text{ in ml}}{1000}$
- Molarity for dilute solutions : $M_1V_1 = M_2V_2$
- For neutralisation reactions : $\frac{M_1V_1}{n_1} = \frac{M_2V_2}{n_2}$
- For mixed solution molarity

$$= \frac{M_1V_1 + M_2V_2 + \dots}{V_1 + V_2 + \dots}$$

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An Introduction to the Principles of Physical Chemistry from the Standpoint of Modern Atomistics and Thermodynamics Edward Wight Washburn,1921

Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh,2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

Solutions Manual to Accompany General Chemistry with Qualitative Analysis, Second Edition Ralph H. Petrucci,Robert K. Wismer,1987

Study Guide with Solutions for Organic Chemistry, Third Edition Ralph J. Fessenden,Joan S. Fessenden,1986 Provides a set of additional drill problems chapter by chapter discussions and supplemental instructional material to help students master organic chemistry problem solving techniques

The Chemical News ,1872

Watts' Dictionary of Chemistry Henry Watts,1892

Metal Industry ,1914

Metal Finishing ,1914

Aliphatic series Victor von Richter,1905

The Chemical News and Journal of Physical Science ,1902

Watts' Dictionary of Chemistry, Revised and Entirely Rewritten Henry Watts,1906

Transactions of the Pharmaceutical Meetings ,1921

The Pharmaceutical Journal and Pharmacist ,1921

Prentice Hall Science Explorer Michael J. Padilla,2002

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