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Civil Engineering Rcc Design

RS Peters



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web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} \text{aq}$ $2\text{PO}_4^{3-} \text{aq}$ $\text{Ba}_3\text{PO}_4\text{2s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{h}^+ \text{aq} + \text{oh}^- \text{aq} \rightarrow \text{h}_2\text{O} \text{l}$ the spectator ions in $\text{kOH} \text{aq}$ and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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web the $\text{Ag}^+ \text{aq}$ and $\text{Cl}^- \text{aq}$ ions become $\text{AgCl} \text{s}$ but the $\text{Na}^+ \text{aq}$ ions and the $\text{NO}_3^- \text{aq}$ ions stay as $\text{Na}^+ \text{aq}$ ions and $\text{NO}_3^- \text{aq}$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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