

Stoichiometry: Mass – Mole Worksheet

Solve the following stoichiometry problems and type in the correct answer with units.
 (for example: 3.45 mol or 4.62 g) Round the masses from the periodic table to the nearest whole number. Round all answers to 2 decimal places.

1. Based on the following chemical equation, please answer the following questions



- a. Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₂.

Answer: _____ $\frac{\text{g FeCl}_2}{\text{g FeCl}_2} \left| \frac{1 \text{ mol FeCl}_2}{\text{g FeCl}_2} \right| \frac{\text{mol NaCl}}{\text{mol FeCl}_2} =$

- b. Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

Answer: _____ $\frac{\text{g Na}_2\text{SO}_4}{\text{g Na}_2\text{SO}_4} \left| \frac{1 \text{ mol Na}_2\text{SO}_4}{\text{g Na}_2\text{SO}_4} \right| \frac{\text{mol Fe}_2(\text{SO}_4)_3}{\text{mol Na}_2\text{SO}_4} \left| \frac{\text{g Fe}_2(\text{SO}_4)_3}{1 \text{ mol Fe}_2(\text{SO}_4)_3} \right| =$

- c. Calculate the amount in grams of FeCl₂ produced with 95g of Fe(SO₄)₂.

Answer: _____ $\frac{\text{g Fe (SO}_4)_2}{\text{g Fe (SO}_4)_2} \left| \frac{1 \text{ mol Fe (SO}_4)_2}{\text{g Fe (SO}_4)_2} \right| \frac{\text{mol FeCl}_2}{\text{mol Fe (SO}_4)_2} \left| \frac{\text{g FeCl}_2}{1 \text{ mol FeCl}_2} \right| =$

- d. Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl.

Answer: _____ $\frac{\text{g NaCl}}{\text{g NaCl}} \left| \frac{1 \text{ mol NaCl}}{\text{g NaCl}} \right| \frac{\text{mol Fe}_2(\text{SO}_4)_3}{\text{mol NaCl}} =$

2. How many moles of CaCO₃ can be produced if we make 12.3 g of Ca(HCO₃)₂ reacts with CaCO₃?



Answer: _____ $\frac{\text{g Ca (HCO}_3)_2}{\text{g Ca (HCO}_3)_2} \left| \frac{1 \text{ mol Ca (HCO}_3)_2}{\text{g Ca (HCO}_3)_2} \right| \frac{\text{mol CaCO}_3}{\text{mol Ca (HCO}_3)_2} =$

Moles And Stoichiometry Packet Answers

Sunil Tanna



Moles And Stoichiometry Packet Answers:

Improving Student Comprehension of Stoichiometric Concepts Connie Lynn Bannick Kemner, 2007 *Forensics in Chemistry* Sara McCubbins, Angela Codron, 2012 Forensics seems to have the unique ability to maintain student interest and promote content learning I still have students approach me from past years and ask about the forensics case and specific characters from the story I have never had a student come back to me and comment on that unit with the multiple choice test at the end from the Introduction to Forensics in Chemistry The Murder of Kirsten K How did Kirsten K's body wind up at the bottom of a lake and what do wedding cake ingredients soil samples radioactive decay bone age blood stains bullet matching and drug lab evidence reveal about whodunit These mysteries are at the core of this teacher resource book which meets the unique needs of high school chemistry classes in a highly memorable way The book makes forensic evidence the foundation of a series of eight hands on week long labs As you weave the labs throughout the year and students solve the case the narrative provides vivid lessons in why chemistry concepts are relevant and how they connect All chapters include case information specific to each performance assessment and highlight the related national standards and chemistry content Chapters provide Teacher guides to help you set up Student performance assessments A suspect file to introduce the characters and new information about their relationships to the case Samples of student work that has been previously assessed and that serves as an answer key for you Grading rubrics Using *Forensics in Chemistry* as your guide you will gain the confidence to use inquiry based strategies and performance based assessments with a complex chemistry curriculum Your students may gain an interest in chemistry that rivals their fascination with *Bones* and *CSI* *Solving Problems in Chemistry* Rod O'Connor, Charles Mickey, Alton Hassell, 1977 [You Can Do Chemistry](#) Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018 **Mole Concepts and Stoichiometry** Dharshi Bopegedera, 2017

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[Bite-Sized Chemistry Calculations](#) John Lambert, 2024-07-10. Bite-Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry, focusing on a few topics at a time to facilitate comprehension. Written by an experienced chemistry educator, each book in the series has been tailored to fully meet the needs of students at all levels, especially those taking college-level general chemistry courses, as well as those following various O-level curricula worldwide. This part of the series explores the different types of problems and calculations encountered in mass, the mole, and stoichiometry, including the determination of formulae of ionic compounds, relative formula masses, mass and percent compositions of compounds, all aspects of mole calculations, empirical and molecular formulae calculations based on chemical equations, limiting reagents, gas stoichiometry, and percent yield. The series is packed with many salient features that are meant to facilitate both teaching and learning. Some of these include helpful explanations, many examples, alternative ways to solve problems, plenty of practice questions, complete answers, and appendices. With this book, you will be well prepared for your exams and boost your performance.

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Answers to Practice Problems

Appendices

Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995. The Study Guide reflects the unique problem-solving approach taken by the Chemical Principles text. The new edition of the Study Guide includes many new worked-out examples.

5 Steps Chemistry Workbook Series Book 2: Mighty Mole Concepts Julie C. Gilbert, 2022-03-07. Chemistry moles got you down? Mole concepts is a challenging unit because there are a lot of different topics. Whether you're a teacher looking for easy worksheets to borrow or a student wanting more practice, I've got something for you. Inside, you'll find:

- Descriptions for each of the major mole concepts topics
- 1 worksheet covering formula mass and molar mass calculations
- 4 worksheets covering various mole conversion topics
- 3 worksheets covering percent calculations
- 4 worksheets covering empirical, molecular, and moles of hydrates calculations
- 2 mixed moles self-tests with answer keys

This is a companion workbook for the *5 Steps to Surviving Chemistry* book and the *5 Steps Chemistry Workbook Series Book 1: Stoichiometry*. However, you do not need to have read those books to find this workbook useful.

Teaching

the Mole Concept in Stoichiometry Timothy Scot Albrecht, California State Polytechnic University, Pomona. College of Education & Integrative Studies, 2009

Calculating in Chemistry Cosmas C Mbuji, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities. Students in higher courses may find it a helpful revision and enhance their clarity. The book completely discusses the key topics in basic stoichiometry including the mole concept, reacting quantities and empirical and molecular formulas. It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance. This is particularly useful to a beginner. Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems. The chapter also describes and applies the concept of a limiting reagent. Two methods of determining a limiting reagent are explained and illustrated. The concept of limiting reagent is extended to reacting volumes of gases with a short cut method. A short cut method for solving reacting quantities involving masses and volumes of gases is also given. Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas. A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner. In chapter 4 we discuss percentage yield and purity. A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances. A student will find these examples helpful in relating different formulas for percentage purity. The last chapter introduces a graphical method for reacting quantities. The method may provide a new way of looking at chemical reactions. Examples are given to illustrate the method including how it can be used to determine limiting reagents. It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course. Teachers may also find this book a good resource for their lessons in stoichiometry.

Unveiling the Magic of Words: A Review of "**Moles And Stoichiometry Packet Answers**"

In some sort of defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Moles And Stoichiometry Packet Answers**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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