

Practice with the Gas Laws with Answers

All temperatures must be in **KELVIN** STP = **0 °C and 1.0 atm**

Converting one unit of pressure to another

1. Change 758. mm Hg to Torr

$$758 \text{ mm Hg} \times \frac{760 \text{ Torr}}{760 \text{ mm Hg}} = 758 \text{ Torr}$$

2. Change 771.3 mm Hg to atm

$$771.3 \text{ mm Hg} \times \frac{1 \text{ atm}}{760 \text{ mm Hg}} = 1.015 \text{ atm}$$

3. Change 95.6 kPa to mm Hg

$$95.6 \text{ kPa} \times \frac{760 \text{ mm Hg}}{101.325 \text{ kPa}} = 717 \text{ mm Hg}$$

4. Is the pressure of the gas sample greater or less than the pressure of the atmosphere? Explain.

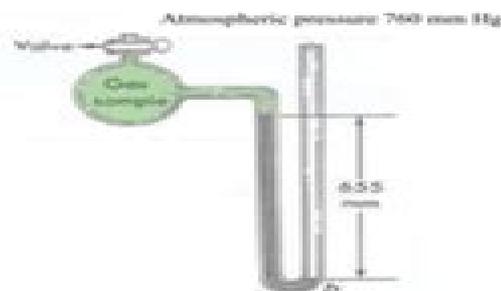
Less, the mercury line is higher on that side

What is the pressure of the gas in this manometer (in mm Hg)?

$$\begin{array}{r} 760 \text{ mm Hg} \\ -655 \text{ mm Hg} \\ \hline 105 \text{ mm Hg} \end{array}$$

Convert that pressure to kPa.

$$105 \text{ mm Hg} \times \frac{101.325 \text{ kPa}}{760 \text{ mm Hg}} = 14.0 \text{ kPa}$$



Calculating Density of a Gas at STP

1. What is the density of dinitrogen tetroxide at STP?

$$\text{N}_2\text{O}_4 \quad \frac{92.01 \text{ g}}{22.4 \text{ L}} = 4.108 \text{ g/L}$$

2. What is the density of sulfur trioxide at STP?

$$\text{SO}_3 \quad \frac{80.06 \text{ g}}{22.4 \text{ L}} = 3.574 \text{ g/L}$$

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Table of Contents Gas Laws Practice Problems With Solutions

1. Understanding the eBook Gas Laws Practice Problems With Solutions
 - The Rise of Digital Reading Gas Laws Practice Problems With Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Practice Problems With Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Practice Problems With Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Practice Problems With Solutions
 - Personalized Recommendations
 - Gas Laws Practice Problems With Solutions User Reviews and Ratings
 - Gas Laws Practice Problems With Solutions and Bestseller Lists
5. Accessing Gas Laws Practice Problems With Solutions Free and Paid eBooks
 - Gas Laws Practice Problems With Solutions Public Domain eBooks
 - Gas Laws Practice Problems With Solutions eBook Subscription Services
 - Gas Laws Practice Problems With Solutions Budget-Friendly Options

6. Navigating Gas Laws Practice Problems With Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Laws Practice Problems With Solutions Compatibility with Devices
 - Gas Laws Practice Problems With Solutions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws Practice Problems With Solutions
 - Highlighting and Note-Taking Gas Laws Practice Problems With Solutions
 - Interactive Elements Gas Laws Practice Problems With Solutions
8. Staying Engaged with Gas Laws Practice Problems With Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Practice Problems With Solutions
9. Balancing eBooks and Physical Books Gas Laws Practice Problems With Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Practice Problems With Solutions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws Practice Problems With Solutions
 - Setting Reading Goals Gas Laws Practice Problems With Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Laws Practice Problems With Solutions
 - Fact-Checking eBook Content of Gas Laws Practice Problems With Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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