

Name _____

Multiple Choice (10 points each)

Identify the letter of the choice that best completes the statement or answers the question.

- E 1. For a certain process at 27°C, $\Delta G = +210.6$ kJ and $\Delta H = -168.2$ kJ. What is the entropy change for this process at this temperature? Express your answer in the form, $\Delta S =$ _____ J/K.
- a. +628.3 J/K
b. -141.3 J/K
c. +141.3 J/K
d. 1.26×10^3 J/K
e. -1.26×10^3 J/K

Answer: $\Delta G = \Delta H - T\Delta S$ $T = 27^\circ\text{C} + 273 = 300\text{K}$

$$\Delta S = \frac{\Delta H - \Delta G}{T} = \frac{(-168.2\text{kJ}) - (+210.6\text{kJ})}{300\text{K}} = -\frac{1.26\text{kJ}}{\text{K}}$$

$$\left(\frac{-1.26\text{kJ}}{\text{K}}\right)\left(\frac{10^3\text{J}}{\text{kJ}}\right) = \boxed{\frac{-1.26 \times 10^3\text{J}}{\text{K}}}$$

- A 2. Which response includes all of the following substances that have $\Delta H_f^\circ = 0$, and no other substances?

- I. HCl(g)
II. Na(s)
III. HCl(aq)
IV. F₂(g)

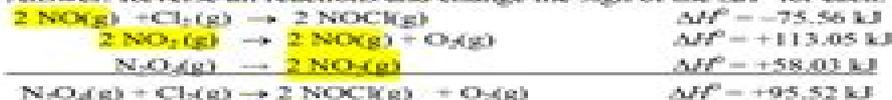
- a. II and IV
b. II
c. I and II
d. I, II, III, and IV
e. I, II, and IV

Answer: All elements in their natural standard state have a $\Delta H_f^\circ = 0$.

- A 3. From the following data at 25°C,

Calculate ΔH° at 25°C for the reaction below.

- a. +95.52 kJ
b. +299 kJ
c. -186.8 kJ
d. -85.52 kJ
e. -246.65 kJ

Answer: Reverse all reactions and change the sign of the ΔH° for each.

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Table of Contents Thermodynamics Problems And Answers

1. Understanding the eBook Thermodynamics Problems And Answers
 - The Rise of Digital Reading Thermodynamics Problems And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermodynamics Problems And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamics Problems And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermodynamics Problems And Answers
 - Personalized Recommendations
 - Thermodynamics Problems And Answers User Reviews and Ratings
 - Thermodynamics Problems And Answers and Bestseller Lists
5. Accessing Thermodynamics Problems And Answers Free and Paid eBooks

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 - Thermodynamics Problems And Answers eBook Subscription Services
 - Thermodynamics Problems And Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Thermodynamics Problems And Answers Compatibility with Devices
 - Thermodynamics Problems And Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermodynamics Problems And Answers
 - Highlighting and Note-Taking Thermodynamics Problems And Answers
 - Interactive Elements Thermodynamics Problems And Answers
 8. Staying Engaged with Thermodynamics Problems And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermodynamics Problems And Answers
 9. Balancing eBooks and Physical Books Thermodynamics Problems And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamics Problems And Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Thermodynamics Problems And Answers
 - Setting Reading Goals Thermodynamics Problems And Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Thermodynamics Problems And Answers
 - Fact-Checking eBook Content of Thermodynamics Problems And Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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