

Name _____

Solution Stoichiometry Worksheet

Solve the following solutions Stoichiometry problems:

1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate?



$$\frac{0.150 \text{ L AgNO}_3}{1 \text{ L}} \times \frac{0.500 \text{ moles AgNO}_3}{1 \text{ L}} \times \frac{1 \text{ moles Ag}_2\text{CrO}_4}{2 \text{ moles AgNO}_3} \times \frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4} = 12.4 \text{ g Ag}_2\text{CrO}_4$$

$$\frac{0.100 \text{ L K}_2\text{CrO}_4}{1 \text{ L}} \times \frac{0.400 \text{ moles K}_2\text{CrO}_4}{1 \text{ L}} \times \frac{1 \text{ moles Ag}_2\text{CrO}_4}{1 \text{ moles K}_2\text{CrO}_4} \times \frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4} = 13.3 \text{ g Ag}_2\text{CrO}_4$$

2. How many mL of 0.280 M barium nitrate are required to precipitate as barium sulfate all the sulfate ions from 25.0 mL of 0.350 M aluminum sulfate? (93.8 mL barium nitrate)



$$\frac{0.0250 \text{ L Al}_2(\text{SO}_4)_3}{1 \text{ L}} \times \frac{0.350 \text{ moles Al}_2(\text{SO}_4)_3}{1 \text{ L}} \times \frac{3 \text{ moles Ba(NO}_3)_2}{1 \text{ moles Al}_2(\text{SO}_4)_3} \times \frac{1 \text{ L}}{0.280 \text{ moles Ba(NO}_3)_2} = 0.0938 \text{ L Ba(NO}_3)_2$$

3. 25.0 mL of 0.350 M NaOH are added to 45.0 mL of 0.125 M copper (II) sulfate. How many grams of copper (II) hydroxide will precipitate?



$$\frac{0.0250 \text{ L NaOH}}{1 \text{ L NaOH}} \times \frac{0.350 \text{ moles NaOH}}{1 \text{ L NaOH}} \times \frac{1 \text{ moles Cu(OH)}_2}{2 \text{ moles NaOH}} \times \frac{97.57 \text{ g Cu(OH)}_2}{1 \text{ mole Cu(OH)}_2} = 0.427 \text{ g Cu(OH)}_2$$

$$\frac{0.0450 \text{ L CuSO}_4}{1 \text{ L NaOH}} \times \frac{0.125 \text{ moles CuSO}_4}{1 \text{ L NaOH}} \times \frac{1 \text{ moles Cu(OH)}_2}{1 \text{ moles CuSO}_4} \times \frac{97.57 \text{ g Cu(OH)}_2}{1 \text{ mole Cu(OH)}_2} = 0.549 \text{ g Cu(OH)}_2$$

4. What volume of 0.415 M silver nitrate will be required to precipitate as silver bromide all the bromide ion in 35.0 mL of 0.128 M calcium bromide?



$$\frac{0.0350 \text{ L CaBr}_2}{1 \text{ L CaBr}_2} \times \frac{0.128 \text{ moles CaBr}_2}{1 \text{ L CaBr}_2} \times \frac{2 \text{ moles AgNO}_3}{1 \text{ moles CaBr}_2} \times \frac{1 \text{ L AgNO}_3}{0.415 \text{ mole AgNO}_3} = 0.0216 \text{ L AgNO}_3$$

5. What volume of 0.496 M HCl is required to neutralize 20.0 mL of 0.809 M sodium hydroxide?



$$\frac{0.0200 \text{ L NaOH}}{1 \text{ L NaOH}} \times \frac{0.809 \text{ mole NaOH}}{1 \text{ L NaOH}} \times \frac{1 \text{ mole HCl}}{1 \text{ mole NaOH}} \times \frac{1 \text{ L HCl}}{0.496 \text{ mole HCl}} = 0.0326 \text{ L HCl}$$

Solution Stoichiometry Worksheet Answer Key

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Solution Stoichiometry Worksheet Answer Key:

Solution Stoichiometry Worksheet Answer Key Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Solution Stoichiometry Worksheet Answer Key**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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