

Electrolytes Worksheet

1. What is an electrolyte? Give an example.

substance that, when dissolved in water, conducts electricity

2. Classify the following substances by type of electrolyte (acid, base or salt).

KOH	<u>B</u>	Ba(NO ₃) ₂	<u>S</u>	KF	<u>S</u>
H ₂ SO ₃	<u>A</u>	HNO ₃	<u>A</u>	Na ₂ CO ₃	<u>S</u>
Mg(OH) ₂	<u>B</u>	NH ₄ OH	<u>B</u>	Fe(OH) ₃	<u>B</u>
HCl	<u>A</u>	MgCl ₂	<u>S</u>	Ca(OH) ₂	<u>B</u>

3. Write the equation for the electrolytic dissociation of the following compounds:

HCl $HCl \xrightarrow{H_2O} H^+(aq) + Cl^-(aq)$

NaOH $NaOH \xrightarrow{H_2O} Na^+(aq) + OH^-(aq)$

H₂SO₄ $H_2SO_4 \xrightarrow{H_2O} 2H^+(aq) + SO_4^{2-}(aq)$

Mg(OH)₂ $Mg(OH)_2 \xrightarrow{H_2O} Mg^{2+}(aq) + 2OH^-(aq)$

BeCl₂ $BeCl_2 \xrightarrow{H_2O} Be^{2+}(aq) + 2Cl^-(aq)$

4. To check the electrical conductivity of certain substances, a student used a conductivity apparatus equipped with a light bulb. Her observations are listed in the following table. Which one of the following groups of substances contains only electrolytes?

Substances	Observations
HCl ✓	Bright light
CH ₃ OH	No light
MgCl ₂ ✓	Faint light
NaOH ✓	Bright light
CH ₃ COOH ✓	Faint light
CCl ₄	No light

- A) CH₃OH and CCl₄ C) CH₃OH, NaOH and CH₃COOH
- B) HCl, MgCl₂ and CCl₄ D) HCl, MgCl₂, NaOH and CH₃COOH

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Table of Contents Electrolytes Worksheet Answer Key

1. Understanding the eBook Electrolytes Worksheet Answer Key
 - The Rise of Digital Reading Electrolytes Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Electrolytes Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electrolytes Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electrolytes Worksheet Answer Key
 - Personalized Recommendations
 - Electrolytes Worksheet Answer Key User Reviews and Ratings
 - Electrolytes Worksheet Answer Key and Bestseller Lists
5. Accessing Electrolytes Worksheet Answer Key Free and Paid eBooks
 - Electrolytes Worksheet Answer Key Public Domain eBooks
 - Electrolytes Worksheet Answer Key eBook Subscription Services
 - Electrolytes Worksheet Answer Key Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Electrolytes Worksheet Answer Key Compatibility with Devices
 - Electrolytes Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electrolytes Worksheet Answer Key
 - Highlighting and Note-Taking Electrolytes Worksheet Answer Key
 - Interactive Elements Electrolytes Worksheet Answer Key
 8. Staying Engaged with Electrolytes Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electrolytes Worksheet Answer Key
 9. Balancing eBooks and Physical Books Electrolytes Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electrolytes Worksheet Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Electrolytes Worksheet Answer Key
 - Setting Reading Goals Electrolytes Worksheet Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Electrolytes Worksheet Answer Key
 - Fact-Checking eBook Content of Electrolytes Worksheet Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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