

Electrochemistry

Q1 Calculate the Emf of the following cell at 298K
 $\text{Fe (s)} | \text{Fe}^{2+} (0.01\text{M}) || \text{H}^+ (1\text{M}) | \text{H}_2(\text{g}) (1\text{bar}), \text{Pt}$
 $E^\circ_{\text{cell}} = 0.44\text{V}$

2023
(2 Marks)



Nernst equation

$$E_{\text{cell}} = E^\circ_{\text{cell}} - \frac{0.059}{n} \log \frac{[\text{Fe}^{2+}] [\text{H}_2]}{[\text{H}^+]^2}$$
$$E_{\text{cell}} = 0.44 - \frac{0.059}{2} \log \frac{[10^{-2}] [1]}{[1]^2}$$
$$E_{\text{cell}} = 0.44 - \frac{0.059}{2} \log 10^{-2}$$
$$E_{\text{cell}} = 0.44 + \frac{0.059}{2} \times 2$$
$$E_{\text{cell}} = \underline{0.499\text{V}}$$

Electrochemistry Problems And Solutions

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Electrochemistry Problems And Solutions:

Unveiling the Power of Verbal Beauty: An Mental Sojourn through **Electrochemistry Problems And Solutions**

In a global inundated with screens and the cacophony of instant connection, the profound power and psychological resonance of verbal artistry usually fade in to obscurity, eclipsed by the constant assault of sound and distractions. Yet, located within the musical pages of **Electrochemistry Problems And Solutions**, a fascinating work of fictional splendor that pulses with raw thoughts, lies an unforgettable journey waiting to be embarked upon. Written by way of a virtuoso wordsmith, this exciting opus guides viewers on a mental odyssey, delicately exposing the latent possible and profound influence embedded within the complex internet of language. Within the heart-wrenching expanse of the evocative evaluation, we can embark upon an introspective exploration of the book is central themes, dissect its fascinating publishing type, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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