

GRAVIMETRIC ANALYSIS PROBLEMS - EXERCISES IN STOICHIOMETRY

1. In the analysis of 0.7011 g of an impure chloride containing sample, 0.9805 g of AgCl were precipitated. What is the percentage by mass chloride in the sample?

$$\begin{aligned}m_{\text{sample}} &= 0.7011 \text{ g} \\m_{\text{AgCl}} &= 0.9805 \text{ g} \\m_{\text{Cl}} &= 0.9805 \text{ g} \times \frac{35.453 \text{ g/mol}}{143.32 \text{ g/mol}} \\&= 0.2425 \text{ g Cl} \\ \% \text{Cl} &= \frac{0.2425 \text{ g Cl}}{0.7011 \text{ g}} \times 100 = 34.59\% \text{Cl}\end{aligned}$$

2. A 0.4054 g solid organic sample containing covalently bound bromide and no other halogens was placed in a porcelain crucible with about one gram of fresh sodium metal. In a process known as *sodium fusion*, the mixture was heated in a furnace to 450°C which charred and vaporized the organic portion of the molecule and converted the covalent bromide into sodium bromide. Excess sodium was decomposed by adding small portions of water which also dissolved the sodium bromide. The clear solution was quantitatively transferred to a beaker, acidified with dilute nitric acid, and diluted to about 50 mL. A 6 mL quantity of 0.1 M AgNO₃ was added to the solution and the mixture heated to about 60°C for an hour to age and digest the precipitate. After filtering, the mass of the silver bromide produced was determined to be 37.8 mg. What is the percentage by mass bromine in the organic compound?

$$\begin{aligned}m_{\text{sample}} &= 0.4054 \text{ g} = 405.4 \text{ mg} \\m_{\text{AgBr}} &= 37.8 \text{ mg} \\ \% \text{Br} &= \frac{37.8 \text{ mg} \times \frac{79.904 \text{ g/mol}}{187.77 \text{ g/mol}}}{405.4 \text{ mg}} \times 100 \\&= 3.97\% \text{ Br}\end{aligned}$$

Gravimetric Analysis Problems Exercises In Stoichiometry

John Charles Olsen



Gravimetric Analysis Problems Exercises In Stoichiometry:

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Library of Congress Subject Headings Library of Congress, Library of Congress. Subject Cataloging Division, 1980
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Hillsdale College Hillsdale College, 1921
Quantitative Chemical Analysis Daniel C. Harris, 2010-04-30
QCA is the bestselling textbook of choice for analytical chemistry It offers a modern portrait of the techniques of chemical analysis backed by a wealth of real world applications This edition features new coverage of spectroscopy and statistics new pedagogy and enhanced lecturer support
A Textbook of Quantitative Analysis Andrew Patterson, 1952
A Text-book of Quantitative Chemical Analysis by Gravimetric, Electrolytic, Volumetric and Gasometric Methods John Charles Olsen, 1910
Year Book, 1914
Library of Congress Subject Headings Library of Congress. Cataloging Policy and Support Office, 2013
Problem Solving in Analytical Chemistry Denise Rafferty, Sara Sleigh, 1998
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The Journal of Education, 1909

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