

Subject	Topic	Test - KAR	Date
C + M + P	Combined Test	JEE Adv - 11 - CT JALL20150928	28 th Sep 2015

C+M+P Key Answers:

1. a	2. d	3. b	4. d	5. a	6. c	7. 5	8. 3	9. 5	10. 5
11. 5	12. b	13. b	14. c	15. b	16. a	17. d	18.	19.	20. a
21. d	22. c	23. a	24. b	25. c	26. 1	27. 3	28. 3	29. 1	30. 4
31. c	32. a	33. b	34. c	35. b	36. a	37.	38.	39. c	40. a
41. c	42. c	43. a	44. b	45. 4	46. 1	47. 5	48. 1	49. 4	50. b
51. c	52. d	53. b	54. d	55. b	56.	57.			

18. A - p, B - p,q, C - r, D - p,s ;

37. A - q, B - r, C - s, D - s,p ;

56. A - s, B - r, C - p, D - q ;

19. A - q, B - r, C - s, D - p ;

38. A - r, B - s, C - p, D - q ;

57. A - r, B - p, C - q, D - t ;

Chemistry Solutions:

- 2-hydroxypropane-1,2,3-tricarboxylic acid
- 2-methylpentane. In 2 position 1 methyl group is there so these compound is having isopropyl group
- 1-Hydroxy - 7 - methyl spiro (2,5) oct - 4 - en - 6 - one
- It is white in colour
- 20ml. dilute unreacted acid soln. required

$$= 32 \text{ mL} \times \frac{N}{10} \text{ NaOH soln.}$$

500 ml. dilute unreacted acid soln. required

$$= \frac{32}{20} \times 500 \text{ mL} \times \frac{N}{10} \text{ NaOH soln.}$$

$$= \frac{32}{20} \times \frac{500}{10} \text{ mL} \times N \text{ NaOH}$$

$$= 80 \text{ mL} \times N \text{ NaOH}$$

$$80 \text{ mL} \times N \text{ NaOH} = 80 \text{ mL} \times N \text{ H}_2\text{SO}_4$$

Acid used for the neutralization of NH_3

$$= (100 - 80) \text{ mL} \times N \text{ H}_2\text{SO}_4$$

$$= 20 \text{ mL} \times N \text{ H}_2\text{SO}_4$$

$$\text{Percentage of nitrogen} = 1.4 \times N_1 \times \frac{V}{W}$$

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1. Understanding the eBook 1542358558 Iit11
 - The Rise of Digital Reading 1542358558 Iit11
 - Advantages of eBooks Over Traditional Books
2. Identifying 1542358558 Iit11
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 1542358558 Iit11
 - User-Friendly Interface
4. Exploring eBook Recommendations from 1542358558 Iit11
 - Personalized Recommendations
 - 1542358558 Iit11 User Reviews and Ratings
 - 1542358558 Iit11 and Bestseller Lists
5. Accessing 1542358558 Iit11 Free and Paid eBooks
 - 1542358558 Iit11 Public Domain eBooks
 - 1542358558 Iit11 eBook Subscription Services
 - 1542358558 Iit11 Budget-Friendly Options
6. Navigating 1542358558 Iit11 eBook Formats
 - ePub, PDF, MOBI, and More
 - 1542358558 Iit11 Compatibility with Devices
 - 1542358558 Iit11 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 1542358558 Iit11
 - Highlighting and Note-Taking 1542358558 Iit11
 - Interactive Elements 1542358558 Iit11
8. Staying Engaged with 1542358558 Iit11

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 1542358558 Iit11
9. Balancing eBooks and Physical Books 1542358558 Iit11
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 1542358558 Iit11
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 1542358558 Iit11
- Setting Reading Goals 1542358558 Iit11
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 1542358558 Iit11
- Fact-Checking eBook Content of 1542358558 Iit11
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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