

1 From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014 (Question 1).

a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	$1.6725 \times 10^{-24}$	$1.6725 \times 10^{-24}$	$0.0009 \times 10^{-24}$

i An atom of hydrogen can be represented as  ${}^1\text{H}$ . Use data from Table 1.1 to calculate the mass of this hydrogen atom.

$$1.6734 \times 10^{-27} \text{ kg}$$

(1 mark)

ii Which one of the following is a fundamental particle that would not be deflected by an electric field?

- A electron  
B neutron  
C proton

Write the correct letter, A, B, or C, on the line below.

B

(1 mark)

b A naturally occurring sample of the element boron has a relative atomic mass of 10.8. In this sample, boron exists as two isotopes,  ${}^{10}\text{B}$  and  ${}^{11}\text{B}$ .

i Calculate the percentage abundance of  ${}^{10}\text{B}$  in this naturally occurring sample of boron.

$$10.8 = (10)(x) + (11)(1-x)$$

$$10.8 = 10x + 11 - 11x$$

$$10.8 - 11 = 10x - 11x$$

$$-0.2 = -x$$

$$x = 0.2 \quad 0.2 \times 100 = 20\%$$

(2 marks)

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## Chapter 11

**Léo Pomar**



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