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**ESPECIFICACIONES DE APRIETE - TORQUE SPECIFICATIONS**  
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**TOYOTA**  
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Year Año	Motor/Engine Displacement CU. IN. [ CC ]	Cylinder Head Bolts	Main Bearing Bolts	Rod Bearing Bolts	Crankshaft Pulley Bolt	Volante Flywheel Bolts	Manifold		Bujias Spark Plugs
							Intake	Exhaust	
1991/92	7M-GE, 7M-GTE 180.3 [2954]	58	75	47	195	54	13	29	13
1991/92	3E-E 88.9 [1457]	9	42	29	112	65	14	35	13
1991/92	4A-FE 97.0 [1587]	44 10	44	36	87	58	14	18	13
1991/92	4A-GE 97.0 [1587]	11	44	12	101	54	20	29	13
1991/92	3S-FE 121.9 [1998]	6	43	36	80	65	14	36	13
1991/92	3S-GTE 121.9 [1998]	6	43	49	80	80	14	38	13
1991/92	5S-FE 132.0 [2164]	6	43	7	80	65	14	36	13
1991/92	2VZ-FE 153.0 [2507]	13	8	7	181	61	13	29	13
1991/92	7M-GE, 7M-GTE 180.3 [2954]	58 10	75	47	195	54	13	29	13

1 New: 65; Used: 63 Lbs pie

2 See text Cabezas/Culatas - automecanico.com

3 3S-GTE: 38

4 29 ft. lbs. and an additional 90° turn Grados

5 M/T: 65 - A/T 61

6 36 ft. lbs. and an additional 90° turn

7 18 ft. lbs. and an additional 90° turn

8 45 ft. lbs. and an additional 90° turn

9 22 ft. lbs., 36 ft. lbs. and an additional 90° turn

10 Torque en sequence, in 3 steps

11 22 ft. lbs. and an additional 90° turn plus  
an additional 90° turn

12 29 ft. lbs. and an additional 90° turn

13 22 ft. lbs. and an additional 90° turn plus  
an additional 90° turn

# Toyota 4y Engine Torque Settings

**SJ Ball**



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monoxide co determine the mass of iodine  $i_2$  which could be produced b if in the above situation only 0 160 moles of iodine  $i_2$

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calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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 $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  c determine the percent yield  
yield actual  $\times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$  theoretical 302.3 g

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c is the limiting reagent 4.33 g of h<sub>2</sub> are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent  
yield theoretical  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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 $\text{CH}_3\text{OH}$  can be produced by the following reaction  $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of co is  
reacted with 8.6 g of h<sub>2</sub>

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  x 1 333 8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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