

## Calculating Average Atomic Mass Worksheet:

- 1) Three isotopes of Silicon occur in nature:

Isotopes of Silicon:	Percent Abundance:	Atomic Mass:
Silicon-28	92.23%	27.97693 amu
Silicon-29	4.68%	28.97649 amu
Silicon-30	3.09%	29.97377 amu

Calculate the average atomic mass for the three isotopes of Silicon.

$$(27.97693 \times 0.9223) + (28.97649 \times 0.0468) + (29.97377 \times 0.0309) = 28.0854 \text{ am}$$

Si-28                      Si-29                      Si-30

- 2) Two isotopes of Rubidium occur naturally:

Isotopes of Rubidium:	Percent Abundance:	Atomic Mass:
Rubidium-85	72.15%	84.9118 amu
Rubidium-87	27.85%	86.9092 amu

Calculate the average atomic mass for the two isotopes of Rubidium.

$$(84.9118 \times 0.7215) + (86.9092 \times 0.2785) = 85.468$$

Rubidium-85                      Rubidium-87

- 3) Only two isotopes of Copper naturally occur:

Isotopes of Copper:	Percent Abundance:	Atomic Mass:
Copper-63	69.17%	62.9396 amu
Copper-65	30.83%	64.9278 amu

Calculate the average atomic mass for the two isotopes of Copper.

$$(62.9396 \times 0.6917) + (64.9278 \times 0.3083) = 63.552$$

Copper-63                      Copper-65

# Calculating Average Atomic Mass Worksheet Answers

**Guillaume Favre**



## **Calculating Average Atomic Mass Worksheet Answers:**

## Unveiling the Magic of Words: A Overview of "**Calculating Average Atomic Mass Worksheet Answers**"

In some sort of defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Calculating Average Atomic Mass Worksheet Answers**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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