

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the $[H_2]$, which will increase the rate, but has no effect on k
 - Due to Arrhenius equation, changing temperature changes the value of k .
 - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/L · s
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

- What is the order with respect to: P_4 0
 H_2 1
- Write the rate law for this reaction. $rate = k[H_2]$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 s^{-1}$**

7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/L · s
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to: SO_2 2
 O_3 0
- Write the rate law for this reaction. $rate = k[SO_2]^2[O_3]^0$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 molL^{-2} s^{-1}$**

8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$

- Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
- Identify any intermediates within the mechanism. **R**
- What is the order with respect to each reactant? **A_2 1st; B_2 1st**

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Solutions Manual for Quanta, Matter and Change Peter Atkins, Julio dePaula, Ron Friedman, 2008-12-15 *US Solutions Manual to Accompany Elements of Physical Chemistry 7e* David Smith, 2017-09-28 The Solutions Manual to Accompany Elements of Physical Chemistry 7th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions

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Applied Mechanics Reviews, 1975 Solutions and Solubilities Michael R. J. Dack, 1975 Physical and Chemical Processes in Gas Dynamics G G Chernyi, Moscow State University S Macheret G Chernyi and S Losev, 2004 A decade ago Russian scientist Staly A Losev assembled top scientists from various Russian institutions at Moscow State University and invited them to pool their knowledge about the changes gases undergo in a variety of aerospace settings from combustion of rocket fuels to hypersonic flow paths to the exteriors of atmospheric reentry vehicles The work grew into the AVAGADRO Laboratory of Moscow State Universitys Institute of Mechanics The models and formulas produced by the team of authors are contained in the second edition of the text book Physical and Chemical Processes in Gas Dynamics The text provides a valuable window into the design philosophies and models that govern former Soviet engineers and scientists Normally these

things are not accessible to U S engineers said Paul Zarchan of MITs Lincoln Laboratory Editor in Chief of AIAs Progress in Astronautics and Aeronautics Series that includes this text Accompanying the book is the Russian produced Chemical Workbench software which is a user friendly Windows based interface for performing calculations based on the models of gas dynamics contained in the book The book and software provide concise exhaustive and clear descriptions of terms notations concepts methods laws and techniques It is invaluable for students and engineers who are determined to more accurately predict the physical and chemical processes in gas and plasma dynamics

Watts' Dictionary of Chemistry, Revised and Entirely Rewritten Henry Watts,1906 Journal of the American Chemical Society American Chemical Society,1914
Proceedings of the Society are included in v 1 59 1879 1937 **Illinois Technograph** ,1956 Transactions of the Faraday Society Faraday Society,1928 **Techniques of Chemistry** ,1975 Electrochemical and Metallurgical Industry Eugene Franz Roeber,Howard Coon Parmelee,1907 **The American Mathematical Monthly** ,1923 Includes section Recent publications **Mathematics of the USSR.** ,1977 Reprint and Circular Series of the National Research Council ,1930 The Journal of Physical Chemistry Wilder Dwight Bancroft,1904 Includes section New Books *The Journal of Physical Chemistry* ,1904 **A Critical Analysis of the Accuracy of Several Numerical Techniques for Combustion Kinetic Rate Equations** Krishnan Radhadrishnan,1993

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