

## Chapter Test B

### Chapter: Chemical Equations and Reactions

Write a correct name (including oxidation state of the metal in ionic formulas) completed with reactant-product arrows and states.

- \_\_\_\_\_ 1. The product of a slightly soluble substance is represented as a double displacement reaction with the other substance as a
  - a. gas
  - b. precipitate**
  - c. insoluble substance
  - d. liquid
- \_\_\_\_\_ 2. The balance chemical equation is represented as follows:
  - a. products**
  - b. reactants
  - c. formation of the products
  - d. number of products
- \_\_\_\_\_ 3. In chemical equation, the coefficient of  $\text{H}_2$  in the reaction  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$  is
  - a. water
  - b. coefficient of water**
  - c. reactant
  - d. coefficient
- \_\_\_\_\_ 4. The following two displacement reactions are given below as they
  - a. chemical reactions are occurring in the following order:  $\text{Fe} > \text{Cu}$
  - b.  $\text{Fe} > \text{Cu}$
  - c.  $\text{Fe} > \text{Cu}$
  - d.  $\text{Fe} > \text{Cu}$**
- \_\_\_\_\_ 5. The following two chemical equations
  - a. indicate the number of moles of  $\text{H}_2$  and  $\text{O}_2$  in the reaction**
  - b. show the number of moles of  $\text{H}_2$  and  $\text{O}_2$  in the reaction
  - c. indicate the number of moles of  $\text{H}_2$  and  $\text{O}_2$  in the reaction
  - d. show the number of moles of  $\text{H}_2$  and  $\text{O}_2$  in the reaction
- \_\_\_\_\_ 6. If  $\text{H}_2$  is shown in the reaction  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ , then
  - a.  $\text{H}_2$  will replace  $\text{O}_2$  in the reaction
  - b.  $\text{H}_2$  will replace  $\text{O}_2$  in the reaction**
  - c.  $\text{H}_2$  will replace  $\text{O}_2$  in the reaction
  - d.  $\text{H}_2$  will replace  $\text{O}_2$  in the reaction

# Chapter 8 Test B

**I. Edward Friedman**



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