

CHAPTER 8 REVIEW

*Chemical Equations and Reactions*

SECTION 1

**SHORT ANSWER** Answer the following questions in the space provided.

1. Match the symbol on the left with its appropriate description on the right.

  d  

(a) A precipitate forms.

  a   ↓

(b) A gas forms.

  b  

(c) A reversible reaction occurs.

  f   (f)

(d) Heat is applied to the reactants.

  e   (aq)

(e) A chemical is dissolved in water.

  c  

(f) A chemical is in the liquid state.

2. Finish balancing the equation:

following



3. In each of the following formulas, write the total number of atoms present.

**12 atoms** \_\_\_\_\_ a.  $4\text{SO}_2$

**16 atoms** \_\_\_\_\_ b.  $3\text{O}_2$

**51 atoms** \_\_\_\_\_ c.  $3\text{Al}_2(\text{SO}_4)_3$

**$3 \times 10^{24}$  atoms** \_\_\_\_\_ d.  $6 \times 10^2 \text{ } 3 \text{ HNO}_3$

4. Convert the following word equation into a balanced chemical equation: aluminum metal + copper(II) fluoride → aluminum fluoride + copper metal



5. One way to test the salinity of a water sample is to add a few drops of silver nitrate solution with a known concentration. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, and sodium nitrate is left in solution. Translate these sentences into a balanced chemical equation.



6. a. Balance the following equation:



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