

Atomic structure and the periodic table

Atomic structure



protons
neutrons
electrons

atoms have no overall charge

particle	relative mass	relative charge
proton	1	+1
neutron	1	0
electron	very small	-1

elements

elements - substances made up of atoms with the same atomic number

isotopes of an element - atoms with the same number of protons but different numbers of neutrons

relative atomic mass - the average mass number for an element

Equations, mixtures and chromatography

chemical formulas and equations

chemical formula - shows the proportion of atoms of each element in a compound

chemical equation - shows the overall change in a reaction

example:

reactants

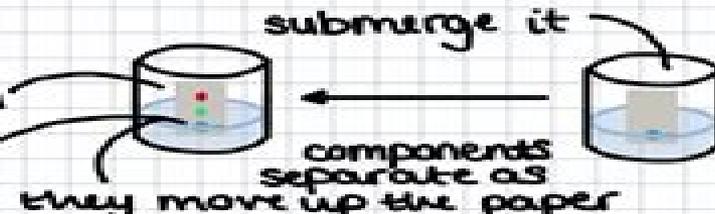
products

word equation: methane + oxygen → carbon dioxide + water

symbol equation: $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

chromatogram

insoluble components stay on the base line



submerge it

components separate as they move up the paper

nuclear symbols

mass number = total number of protons and neutrons in an atom



atomic number = number of protons in an atom

compounds

compound - a substance formed from two or more elements chemically bonded together in fixed proportions

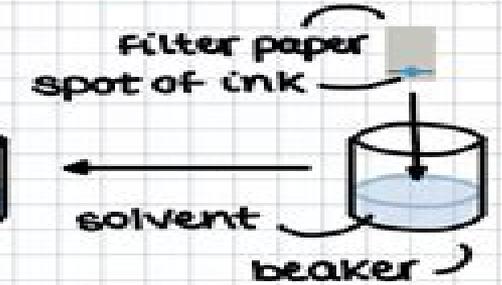
molecule - particle containing two or more non-metal atoms bonded covalently. Molecules can be elements or compounds

mixtures

mixtures - substances made up of different elements or compounds that are not chemically bonded to each other

mixtures can be separated using physical methods

paper chromatography



filter paper

spot of ink

solvent

beaker

Chapter 7 Notes Atomic Structure And Periodicity

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