

Temperature $^{\circ}\text{C}$, K

measure of average KE
of motion of particles

Heat

kJ, kcal (Cal) [1 kcal = 4.184 kJ]

measure of total energy transferred from
an object of high E \rightarrow low E.

Note: A change in T is accompanied by a transfer of heat energy.

Specific heat (c or c_p):

(specific for each substance)

amount of energy required to raise the temperature of 1 g of that substance by 1°C .

change in heat = specific heat \times mass \times change in T

$$q = c_p \times m \times \Delta T$$

Units: heat absorbed or released, in J;

ΔT in $^{\circ}\text{C}$ or K; m in g; c in $\text{J/g}^{\circ}\text{C}$

Molar heat capacity = heat capacity in $\text{J/mol}^{\circ}\text{C}$ instead of $\text{J/g}^{\circ}\text{C}$

Don't forget: the size of $1^{\circ}\text{C} = 1 \text{ K}$, so K can be used in this calculation in place of $^{\circ}\text{C}$.

Examples

- How much heat energy is released to your body when a cup of hot tea containing 200. g of water is cooled from 65.0°C to body temperature, 37.0°C ? (c_p of $\text{H}_2\text{O} = 4.18 \text{ J/g}^{\circ}\text{C}$)
- A 4.50-g nugget of pure gold absorbed 276 J of heat. What was the final temperature of the gold if the initial temperature was 25.0°C ? (c_p of gold = $0.129 \text{ J/g}^{\circ}\text{C}$)
- A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C . In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance?

Chapter 6 Thermochemistry Test

M Walker



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