

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and make use of Le Chatelier's principle and an engineering design process through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{Fe}(\text{OH})^{2+}(\text{aq}) + \text{H}^{+}(\text{aq})$$
4. In part 1, what is the conjugate acid and base? In part 2, what is the conjugate acid and base?
Conjugate acids have an extra proton and base have the other hand but nothing. The conjugate base has no protons, the conjugate acid has one. And conjugate acids have one extra proton.
5. What is the conjugate acid and base? In part 3?
There is a proton on the left side of the reaction so that conjugate acid is on the left.
6. What is the conjugate acid and base? In part 4?
There are no protons on the right side of the reaction so the conjugate acid is on the right.
7. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+} + \text{OH}^{-} \rightleftharpoons \text{FeOH}^{2+}$$
8. Write your own reaction that $[\text{Fe}^{3+}]$ and $[\text{SCN}^{-}]$ will react. The products must be FeSCN^{2+} and water. H^{+} is the product.
$$\text{Fe}^{3+} + \text{SCN}^{-} \rightleftharpoons \text{FeSCN}^{2+} + \text{H}^{+}$$
9. What is a spectator ion?
A spectator ion means in the reaction there are ions that neither participate in the reaction. In equilibrium and do not affect equilibrium.

Investigating Chemical Equilibrium Lab 12a Answer

Annelies Wilder-Smith



Investigating Chemical Equilibrium Lab 12a Answer:

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