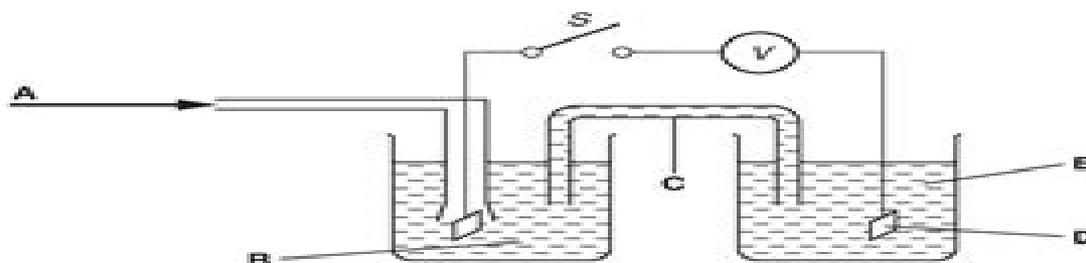


ELECTROCHEMISTRY WS 1

- 1 Chlorine gas and iron(II) ions react together in aqueous solution as follows.



- (a) The following diagram shows the apparatus needed to measure the E_{cell}° for the above reaction.



- (i) In the spaces below, identify what the five letters A – E in the above diagram represent.

A chlorine at $P_{\text{Cl}_2} = 1 \text{ atm}$ and 298 K
 B $1 \text{ mol dm}^{-3} \text{ NaCl}$
 C salt bridge
 D platinum
 E $1 \text{ mol dm}^{-3} \text{ Fe}^{2+} \text{ aq}$ and $1 \text{ mol dm}^{-3} \text{ Fe}^{3+} \text{ aq}$

- (ii) Use the *Data Booklet* to calculate the E_{cell}° for this reaction, and hence decide which direction (left to right, or right to left) electrons would flow through the voltmeter V when switch S is closed.



[7]

Electrochemistry Answers

B Lingard



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Table of Contents Electrochemistry Answers

1. Understanding the eBook Electrochemistry Answers
 - The Rise of Digital Reading Electrochemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electrochemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electrochemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electrochemistry Answers
 - Personalized Recommendations
 - Electrochemistry Answers User Reviews and Ratings
 - Electrochemistry Answers and Bestseller Lists
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 - Highlighting and Note-Taking Electrochemistry Answers
 - Interactive Elements Electrochemistry Answers
 8. Staying Engaged with Electrochemistry Answers
 - Joining Online Reading Communities
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 - Minimizing Distractions
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 11. Cultivating a Reading Routine Electrochemistry Answers
 - Setting Reading Goals Electrochemistry Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Electrochemistry Answers
 - Fact-Checking eBook Content of Electrochemistry Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

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 - Integration of Multimedia Elements
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