

## Introduction 11

# An Electron's Address

Name: \_\_\_\_\_  
 Date: \_\_\_\_\_  
 Class: \_\_\_\_\_

### Introduction: Energy Levels and Sublevels

In addition to the energy level model, Bohr proposed the following model of energy levels. The energy level of the atom can be thought of as the distance

between adjacent energy levels. Energy levels are the sublevels of the energy level. Each energy level contains several sublevels. Draw the following table:

Table 1

Energy Level	Number of sublevels in each energy level
1	1
2	2
3	3
4	4

Draw the table in reverse order as a structural model of the  $1^{\text{st}}$  energy level. Draw the energy level of sublevels in the  $1^{\text{st}}$  energy level.

### Critical Thinking Questions

- How many sublevels exist in the  $1^{\text{st}}$  energy level?  
 The sublevels of the energy level are the first energy level of the atom which is the only one sublevel.
- How many sublevels exist in the  $2^{\text{nd}}$  energy level?  
 Two sublevels exist in the second energy level of the atom which are the s and p sublevels.
- How many sublevels exist in the  $3^{\text{rd}}$  energy level?  
 Three sublevels exist in the third energy level and these are the s, p, and d sublevels.
- How many sublevels would you expect to have for the  $4^{\text{th}}$  energy level?  
 The number of sublevels I would expect to have in the 4th energy level would be 4 sublevels. There would be 4 sublevels which are the s, p, d, and f sublevels.
- How many sublevels would you expect to have for the  $5^{\text{th}}$  energy level?  
 The number of sublevels I would expect to have in the 5th energy level would be 5 sublevels. There would be 5 sublevels which are the s, p, d, and f sublevels.

# Chemquest 11 Answers

**Clemens Wendtner**



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