

Nomenclature Review

Name: _____

1. Give the charge and name for each of the following *oxyanions*.

	Charge	Name
a. NO ₃	-1	nitrate
b. CO ₃	-2	carbonate
c. C ₂ H ₃ O ₂	-1	acetate
d. ClO ₂	-1	chlorite
e. PO ₃	-3	phosphite
f. ClO ₄	-1	perchlorate
g. SO ₄	-2	sulfate
h. NH ₄	+1	ammonium
i. OH	-1	hydroxide
j. CN	-1	cyanoide
k. BrO ₃	-1	bromate
l. C ₂ O ₄	-2	oxalate
m. SO ₃	-2	sulfate
n. MnO ₄	-1	permanganate

2. If A and B represent compounds of an element bonded to a polyatomic ion, give an example of a valid formula for each combination below.

a. AB	b. AB ₂	c. A ₂ B	d. AB ₃	e. A ₂ B ₃
Al(NO ₃) ₃	Ca(NO ₂) ₂	(NH ₄) ₂ S	Al(ClO ₄) ₃	Al ₂ (SO ₄) ₃

3. What is the oxidation number of:

a. Calcium ion?	+2		
b. Sodium ion?	+1		
c. Sulfur ion?	-2		
d. Phosphorus ion?	-3		
e. Carbon ion within Carbon dioxide molecule?	+4	$x + 2(-2) = 0$ $x = 4+$	
f. Phosphorus within the Phosphate polyatomic ion?	+5		$x + 4(-2) = -2$ $x = +6$
g. Chromium within the Chromate polyatomic ion?	+6		
h. Carbon within the Oxalate polyatomic ion?	+3		$x + 4(-2) = -2$ $2x + 4(-2) = -2$ $x = 3+$
i. Zinc ion?	+2		
j. Cadmium ion?	+2		
k. Sulfur within the Hydrogen Sulfite ion (HSO ₃) ⁻	+4	$1 + x - 6 = -1$ $x = 4+$	

4. Give the most common oxidation number/numbers for each of the following groups or families of elements.

a. Alkali Metals	+1
b. Alkaline Earth	+2
c. Noble Gases	0
d. Halogens	-1
e. Group 13	+3
f. Group 15	-3

5. Give the formula for the following hydrocarbons/alcohols:

a. methane	C ₁ H ₄
b. butene	C ₄ H ₈
c. ethanol	C ₂ H ₅ O ₁
d. propyne	C ₃ H ₄

Chemistry Unit 4 Test Review Electron Configuration

Drew H. Wolfe



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