

Unit Two Chemistry Exam

Multiple Choice 20 marks

You need to record your answers in the appropriate section in the ANSWER BOOKLET.

Question 1

When ethanol dissolves in water the bonds formed between the ethanol and water molecules are known as

- A. dispersion forces
- B. ion-dipole attraction
- C. hydrogen bonds
- D. ionisation

Question 2

When NaOH dissociates in water, the equation that describes the process is

- A. $\text{NaOH (s)} \rightarrow \text{Na}^+ (\text{aq}) + \text{OH}^- (\text{aq})$
- B. $\text{NaOH (s)} + \text{H}_2\text{O (l)} \rightarrow \text{NaO}^- (\text{aq}) + \text{H}_2\text{O}^+ (\text{aq})$
- C. $\text{NaOH (s)} + \text{H}_2\text{O (l)} \rightarrow \text{NaOH}_2^+ (\text{aq}) + \text{OH}^- (\text{aq})$
- D. $\text{NaOH (aq)} + \text{HCl (aq)} \rightarrow \text{NaCl (s)} + \text{H}_2\text{O (l)}$

Question 3

How much heat energy is required to raise the temperature of 250 mL of water from 17 °C to 100 °C? Assume that 1 mL of water has a mass of 1 g and that the specific heat capacity of water is $4.2 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$.

- A. 8.7 J
- B. 87,150 KJ
- C. 8.7 KJ
- D. $8.7 \times 10^4 \text{ J}$

Question 4

A solution containing 0.25 g of NaCl dissolved in 50 mL of deionised water will have a concentration of

- A. 5.0 gL^{-1}
- B. 4.3 M
- C. $5.0 \times 10^{-3} \text{ mgL}^{-1}$
- D. $2.5 \times 10^{-4} \text{ molL}$

Question 5

A solution of CuSO_4 has a concentration of 2.50 M. This concentration is equivalent to

- A. 399 %m/v
- B. 39.9 %m/v
- C. 0.250 %m/v
- D. 25.0 m/v

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