

Periodic Trends Worksheet

Using the periodic table and your knowledge of periodic trends, answer the following questions:

1 Circle the best answer that has the:

- | | | | | | | |
|---|----------------------------|----|----|----|----|----|
| a | smallest electron affinity | N | P | As | Sb | Bi |
| b | largest electronegativity | Ga | Ge | As | Sc | |
| c | largest ionization energy | O | S | Sc | Te | Po |
| d | lowest ionization energy | O | S | Se | Te | Po |
| e | smallest atomic radius | F | Cl | Br | I | At |
| f | largest atomic radius | Li | Na | K | Rb | Cs |

2 Low ionization energy is a characteristic of a _____

3 What does atomic radius measure? _____

4 Describe the change in the atomic size across a period and down a group.

5 Which element has the largest atomic radius? _____

6 Explain the main difference between metals and nonmetals.

7 Determine if the listed pairs of elements have similar or different properties.

- | | | | |
|---|----------------------|---|--------------------|
| a | Potassium & Rubidium | b | Calcium & Barium |
| c | Sodium & Chlorine | d | Helium & Krypton |
| e | Phosphorus & Oxygen | f | Lithium & Fluorine |

8 Which element has the largest electronegativity? _____

9 Which group (name) would have the lowest ionization energies?

10 Which group (name) has the highest ionization energies?

Chemistry Practice Test Periodic Trends And Orbitals

**Daniel Barker Murphy, Viateur
Rousseau**



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The Science Teacher, 1995

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