

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## States of Matter

## Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

- The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
- The kinetic molecular theory makes the following assumptions:
  - In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
  - Between gas particles and the walls, there are no significant attractive or repulsive forces between gas particles.
  - Gas particles are in constant and random motion.
  - The collisions between gas particles are elastic, that is, no kinetic energy is lost.
- The kinetic energy of a particle is represented by the equation  $KE = \frac{1}{2}mv^2$ .
- Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- |              |  |
|--------------|--|
| <u>True</u>  | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas.                              |
| <u>True</u>  | 2. The random motion of gas particles causes a gas to expand until it fills its container.                                       |
| <u>False</u> | 3. The density of a gas decreases as it is compressed.   |
| <u>True</u>  | 4. A gas can flow into a space occupied by another gas.  |
| <u>True</u>  | 5. The diffusion of a gas is caused by the random motion of the particles of the gas.  |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles.   |
| <u>True</u>  | 7. During effusion, a gas escapes through a tiny opening into a vacuum.  |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

Study Guide for Content Mastery

Chemistry: Matter and Change • Chapter 13 79

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## CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## Section 13.1 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

- Pressure is defined as force per unit
  - area
  - time
  - mass
  - volume
- What is an instrument designed to measure atmospheric pressure?
  - barometer
  - manometer
  - physoconimeter
  - densimeter
- The height of the liquid in a barometer is affected by all of the following EXCEPT the
  - altitude
  - density of the liquid in the column
  - atmospheric pressure
  - diameter of the column tube.
- The pressure of the gas in a manometer is directly related to which of the following quantities?
  - height of the mercury column in the closed-end arm
  - height of the mercury column in the open-end arm
  - $h + h_0$
  - $h - h_0$
- One atmosphere is equal to a pressure of
  - 76 mm Hg
  - 101.3 kPa
  - 147 psi
  - 760 mm
- The partial pressure of a gas depends on all of the following EXCEPT the
  - concentration of the gas
  - size of the container
  - density of the gas
  - temperature of the gas
- The pressure of a sample of air in a manometer is 101.3 kPa. What is the partial pressure of nitrogen ( $N_2$ ) in the sample if the combined partial pressures of the other gases is 22.4 kPa?
  - 42.4 kPa
  - 78.9 kPa
  - 101.3 kPa
  - 123.7 kPa

Use the figure to answer the following questions.

- What instrument is illustrated in the figure? barometer
- Who invented this instrument? Evangelista Torricelli
- What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
- What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



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# Chemistry Matter And Change Chapter 18 Study Answers

**C Cleary**



## **Chemistry Matter And Change Chapter 18 Study Answers:**

**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement **Instructor's Manual and Solutions to Problems to Accompany**

**Chemistry, a Study of Matter** Alfred Benjamin Garrett, William Thomas Lippincott, Frank Henry Verhoek, 1972

CHEMISTRY SILBERBERG, 2003 Know Your 'O' Level Chemistry - A Study Guide , **Chemistry** James E.

Brady, John R. Holum, 1993-01-21 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples **Chemistry, Student Solutions Manual** James E. Brady, Joel W.

Russell, John R. Holum, 2000-01-26 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of  $\text{C}_2\text{H}_2$  and  $\text{O}_2$  to give  $\text{CO}_2$  and  $\text{H}_2\text{O}$  Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet Chemistry, Study Guide James E. Brady, John R. Holum, 1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to

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**Chemistry** Henry Dorin, 1987 Chemistry Martin Stuart Silberberg, 1996 This general chemistry text centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives

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**Nature** Sir Norman Lockyer, 1909 The

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