

# Chemistry

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(Chapter 1)(Solutions) (InText Questions)  
(Class 12)

## Question 1.11:

Calculate the mass of ascorbic acid (Vitamin C,  $C_6H_8O_6$ ) to be dissolved in 75 g of acetic acid to lower its melting point by  $1.5^\circ\text{C}$ .  $K_f = 3.9 \text{ K kg mol}^{-1}$ .

## Answer 1.11:

Mass of acetic acid,  $w_1 = 75 \text{ g}$

Molar mass of ascorbic acid ( $C_6H_8O_6$ ),  $M_2 = 6 \times 12 + 8 \times 1 + 6 \times 16 = 176 \text{ g mol}^{-1}$

Lowering of melting point,  $\Delta T_f = 1.5 \text{ K}$

We know that:

$$\Delta T_f = \frac{K_f \times 1000 \times w_2}{M_2 \times w_1}$$

$$\Rightarrow w_2 = \frac{\Delta T_f \times M_2 \times w_1}{K_f \times 1000}$$

$$= \frac{0.5 \times 176 \times 75}{3.9 \times 1000}$$

$$= 5.08 \text{ g (approx)}$$

Hence, 5.08 g of ascorbic acid is needed to be dissolved.

## Question 1.12:

Calculate the osmotic pressure in Pascals exerted by a solution prepared by dissolving 1.0 g of polymer of molar mass 185,000 in 450 mL of water at  $37^\circ\text{C}$ .

## Answer 1.12:

It is given that:

Volume of water,  $V = 450 \text{ mL} = 0.45 \text{ L}$

Temperature,  $T = (37 + 273) \text{ K} = 310 \text{ K}$

Number of moles of the polymer,  $n = \frac{1}{185000} \text{ mol}$

We know that:  $\pi = \frac{n}{V} RT$

Osmotic pressure,

$$= \frac{1}{185000} \text{ mol} \times \frac{1}{0.45 \text{ L}} \times 8.314 \times 10^3 \text{ Pa L K}^{-1} \times 310 \text{ K}$$

$$= 30.98 \text{ Pa}$$

$$= 31 \text{ Pa (approximately)}$$

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# Chemistry For Today Answers And Solutions

**Ruth Gavison**



## **Chemistry For Today Answers And Solutions:**

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interception and characterization of key intermediates either as transient species or as protonated deprotonated forms of neutral species by API MS The outstanding features and advantages of ESI MS make it one of the most suitable tools for the fast screening of intermediates directly from solution providing hitherto unavailable chemical information to organic chemists This monograph provides an overview of the mechanisms involved in ESI MS the historical perspectives before looking further in depth at specific reactions and intermediates Written by researchers in the field this book is an unique resource for the understanding of this cutting edge technique

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