

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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Table of Contents Chemistry Chapter 10 The Mole Study Guide Answers

1. Understanding the eBook Chemistry Chapter 10 The Mole Study Guide Answers
 - The Rise of Digital Reading Chemistry Chapter 10 The Mole Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 10 The Mole Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 10 The Mole Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 10 The Mole Study Guide Answers
 - Personalized Recommendations

- Chemistry Chapter 10 The Mole Study Guide Answers User Reviews and Ratings
- Chemistry Chapter 10 The Mole Study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Chapter 10 The Mole Study Guide Answers Free and Paid eBooks
 - Chemistry Chapter 10 The Mole Study Guide Answers Public Domain eBooks
 - Chemistry Chapter 10 The Mole Study Guide Answers eBook Subscription Services
 - Chemistry Chapter 10 The Mole Study Guide Answers Budget-Friendly Options
- 6. Navigating Chemistry Chapter 10 The Mole Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 10 The Mole Study Guide Answers Compatibility with Devices
 - Chemistry Chapter 10 The Mole Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 10 The Mole Study Guide Answers
 - Highlighting and Note-Taking Chemistry Chapter 10 The Mole Study Guide Answers
 - Interactive Elements Chemistry Chapter 10 The Mole Study Guide Answers
- 8. Staying Engaged with Chemistry Chapter 10 The Mole Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 10 The Mole Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 10 The Mole Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 10 The Mole Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 10 The Mole Study Guide Answers
 - Setting Reading Goals Chemistry Chapter 10 The Mole Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 10 The Mole Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Chapter 10 The Mole Study Guide Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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