

SYLLABUS

Mole Concept and Stoichiometry

- (i) Gay-Lussac's Law of Combining Volumes; Avogadro's Law.
Idea of mole – a number just as dozen, a gross; Avogadro's Law – statement and explanation; Gay-Lussac's Law of Combining Volumes – statement and explanation; "the mass of 22.4 litres of any gas at S.T.P. is equal to its molar mass" (Questions will not be set on formal proof but may be taught for clear understanding) – simple calculations based on the molar volume.
- (ii) Refer to the atomicity of hydrogen, oxygen, nitrogen and chlorine (proof not required).
The explanation can be given using equations for the formation of HCl, NH₃ and NO.
- (iii) Relative atomic masses (atomic weight) and relative molecular masses (molecular weights): either H = 1 or ¹²C = 12 will be accepted; molecular mass = 2 × vapour density (formal proof not required). Deduction of simple (empirical) and molecular formula from the percentage composition of a compound; the molar volume of a gas at S.T.P.; simple calculations based on chemical equations; both reacting weight and volumes.
*Idea of relative atomic mass and relative molecular mass – standard H atom, or 1/12th of carbon 12 atom.
 Relating mole and atomic mass, arriving at gram atomic mass and then gram atom; atomic mass is a number dealing with one atom; gram atomic mass is the mass of one mole of atoms.
 Relating mole and molecular mass arriving at gram molecular mass and gram molecule – molecular mass is a number dealing with a molecule, gram molecular mass is the mass of one mole of molecules.
 Molecular mass = 2 × vapour density (questions will not be set on formal proof but may be taught for clear understanding); simple calculations based on the formula.
 Deduction of simple (empirical) and molecular formula from the percentage composition of a compound.*

5A. GAY-LUSSAC'S LAW AND AVOGADRO'S LAW

5.1 INTRODUCTION

You have learnt in Class IX that all gases behave similarly under similar conditions of temperature and pressure as expressed by gas laws.

Gas laws are certain rules applicable to a gas in respect of change in temperature (*T*) or pressure (*P*) or volume (*V*). The change in any one of these affects the other two.

Pressure-Volume Relationship or Boyle's Law

It states that *the volume of a given mass of dry gas is inversely proportional to its pressure at a constant temperature.*

$$P_1 V_1 = P_2 V_2 = k \text{ at constant temperature}$$

Temperature-Volume Relationship or Charles's Law :

It states that *volume of a given mass of a dry gas is directly proportional to its absolute (kelvin) temperature, if the pressure is kept constant.*

OR

The pressure remaining constant, the volume of a given mass of a dry gas increases or decreases by 1/273 of its volume for each 1°C increase or decrease in temperature respectively.

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} = k \text{ at constant pressure}$$

Gas Equation

On combining both the laws we conclude that

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