

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
 potassium-40 undergoes beta decay (electron)  
 uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.  
 Write the electron configuration for Cr.  
 Write the shorthand notation for Bi.



### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- |                             |   |
|-----------------------------|---|
| a. Sulfur hexafluoride      | $\text{SF}_6$   |
| b. Lithium nitride          | $\text{Li}_3\text{N}$   |
| c. Chromium (III) carbonate | $\text{Cr}_2(\text{CO}_3)_3$  |
| d. Tin (II) chloride        | $\text{SnCl}_2$   |
| e. Ammonium acetate         | $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$ ) |
| f. Mercury (I) chloride     | $\text{Hg}_2\text{Cl}_2$  |
| g. Potassium bromate        | $\text{KBrO}_3$   |
| h. Hydrosulfuric acid       | $\text{H}_2\text{S(aq)}$  |
| i. Chloric acid             | $\text{HClO}_3$   |
| j. Sulfurous acid           | $\text{H}_2\text{SO}_3$   |

Name each of the following compounds:

- |                              |                         |
|------------------------------|-------------------------|
| a. $\text{CuSO}_4$           | Copper (II) sulfate     |
| b. $\text{AlF}_3$            | aluminum fluoride       |
| c. HI                        | hydrogen iodide         |
| d. NO                        | nitrogen monoxide       |
| e. $\text{H}_2\text{Se(aq)}$ | hydroselenic acid       |
| f. $\text{HNO}_3$            | nitric acid             |
| g. $\text{HNO}_2$            | nitrous acid            |
| h. $\text{NaHSO}_4$          | sodium hydrogen sulfate |

### Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



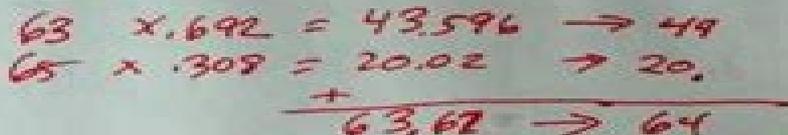
Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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**Peterson's**



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  - The Rise of Digital Reading Chemistry 1 Honors Final Exam Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 1 Honors Final Exam Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry 1 Honors Final Exam Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 1 Honors Final Exam Study Guide
  - Personalized Recommendations
  - Chemistry 1 Honors Final Exam Study Guide User Reviews and Ratings
  - Chemistry 1 Honors Final Exam Study Guide and Bestseller Lists
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  - Chemistry 1 Honors Final Exam Study Guide Enhanced eBook Features
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  - Interactive Elements Chemistry 1 Honors Final Exam Study Guide
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- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry 1 Honors Final Exam Study Guide
9. Balancing eBooks and Physical Books Chemistry 1 Honors Final Exam Study Guide
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry 1 Honors Final Exam Study Guide
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 1 Honors Final Exam Study Guide
- Setting Reading Goals Chemistry 1 Honors Final Exam Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 1 Honors Final Exam Study Guide
- Fact-Checking eBook Content of Chemistry 1 Honors Final Exam Study Guide
  - Distinguishing Credible Sources
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- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

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- Interactive and Gamified eBooks

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