

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the $[H_2]$, which will increase the rate, but has no effect on k
 - Due to Arrhenius equation, changing temperature changes the value of k .
 - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/L · s
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

- What is the order with respect to: P_4 0
 H_2 1
- Write the rate law for this reaction. $rate = k[H_2]$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 s^{-1}$**

7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/L · s
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to: SO_2 2
 O_3 0
- Write the rate law for this reaction. $rate = k[SO_2]^2[O_3]^0$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 molL^{-2} s^{-1}$**

8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$

- Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
- Identify any intermediates within the mechanism. **R**
- What is the order with respect to each reactant? **A_2 : 1st; B_2 : 1st**

Chemical Kinetics Practice Problems And Solutions

Arthur Woolsey Ewell



Chemical Kinetics Practice Problems And Solutions:

Calculations in Chemical Kinetics for Undergraduates Eli Usheunepa Yunana, 2022-06-14 Calculations in Chemical Kinetics for Undergraduates aims to restore passion for problem solving and applied quantitative skills in undergraduate chemistry students Avoiding complicated chemistry jargon and providing hints and step wise explanations in every calculation problem students are able to overcome their fear of handling mathematically applied problems in physical chemistry This solid foundation in their early studies will enable them to connect fundamental theoretical chemistry to real experimental applications as graduates Additional Features Include Contains quantitative problems from popular physical chemistry references Provides step by step explanations are given in every calculation problem Offers hints to certain problems as points to note to enable student comprehension Includes solutions for all questions and exercises This book is a great resource for undergraduate chemistry students however the contents are rich and useful to even the graduate chemist that has passion for applied problems in physical chemistry of reaction Kinetics **(Free Sample) GO TO Objective NEET**

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Schaum's Outlines Problem Solved **A Text-book of Physical Chemistry, Theory and Practice** Arthur Woolsey Ewell, 1909 **College of Engineering Catalogue** Carnegie Institute of Technology. College of Engineering and Science, 1927 Student's Guide to Chemistry; a Modern Introduction David Brooks, 1974 *Cornell University Announcements* Cornell University, 1925 *Organic Chemistry* Paula Yurkanis Bruice, 1998. Written by the author, this student aid features complete step-by-step solutions to all exercises in the text, an essay on electron pushing formalism, etc.

Chemical & Metallurgical Engineering Eugene Franz Roeber, Howard Coon Parmelee, 1921 *Chemical Kinetics and Dynamics* Jeffrey I. Steinfeld, Joseph Salvadore, Francisco, William L. Hase, 1999. This text presents a balanced presentation of the macroscopic view of empirical kinetics and the microscopic molecular viewpoint of chemical dynamics. This second edition includes the latest information as well as new topics such as heterogeneous reactions in atmospheric chemistry, reactant-product imaging, and molecular dynamics of H₂.

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