

9. Evaluate the experiments that led to the conclusion that electrons are negatively charged particles found in all matter.

The deflection toward positively charged plates demonstrated the negatively charged nature of electrons; the fact that changing the type of electrode or the type of gas used in the cathode ray tube did not affect the ray produced led to the conclusion that electrons are present in all matter.

10. Compare the relative charge and mass of each of the subatomic particles.

Particle	Relative charge	Relative mass
Electron	-1	1/1840
Proton	+1	1
Neutron	0	-1

11. **Calculate** What is the difference expressed in kilograms between the mass of a proton and the mass of an electron?

$$1.673 \times 10^{-24} \text{ g} - 9.11 \times 10^{-28} \text{ g} = 1.672 \times 10^{-24} \text{ g} \times 1 \text{ kg}/1000 \text{ g} = 1.672 \times 10^{-27} \text{ kg}$$

Section 4.3 How Atoms Differ

pages 115–121

Practice Problems

pages 116–121

12. How many protons and electrons are in each atom?
- radon, 86 protons and 86 electrons
 - magnesium, 12 protons and 12 electrons
13. An atom of an element contains 66 electrons. Which element is it?
dysprosium
14. An atom of an element contains 14 protons. Which element is it?
silicon

15. **Challenge** Do the atoms shown in the figure to the right have the same atomic number?



Yes, 9

16. Determine the number of protons, electrons, and neutrons for isotopes b through f, in the table above. Name each isotope, and write its symbol.

	Protons and Electrons	Neutrons	Isotope	Symbol
b.	20	26	calcium-46	$^{46}_{20}\text{Ca}$
c.	8	9	oxygen-17	$^{17}_8\text{O}$
d.	26	31	iron-57	$^{57}_{26}\text{Fe}$
e.	30	34	zinc-64	$^{64}_{30}\text{Zn}$
f.	80	124	mercury-204	$^{204}_{80}\text{Hg}$

17. **Challenge** An atom has a mass number of 55. Its number of neutrons is the sum of its atomic number and 5. How many protons, neutrons, and electrons does this atom have? What is the identity of this atom?

25 protons, 25 electrons, 30 neutrons, manganese

18. Boron (B) has two naturally occurring isotopes: boron-10 (abundance = 19.8%, mass = 10.013 amu), boron-11 (abundance = 80.2%, mass = 11.009 amu). Calculate the atomic mass of boron.

$$\text{B-10 } 10.013 \text{ amu} \times 0.198 = 1.98 \text{ amu}$$

$$\text{B-11 } 11.009 \text{ amu} \times 0.802 = 8.83 \text{ amu}$$

$$\text{Atomic mass } 1.98 \text{ amu} + 8.83 \text{ amu} = 10.81 \text{ amu}$$

19. **Challenge** Nitrogen has two naturally occurring isotopes, N-14 and N-15. Its atomic mass is 14.007. Which isotope is more abundant? Explain your answer.

N-14 is more abundant because the atomic mass is closer to 14 than 15.

Chapter 4 Atomic Structure Workbook Answers

John Allen



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Learning Chemistry 7 Solution Book (Year 2023-24), 2024-01-02 *Learning Elementary Chemistry Class 7 Teacher Resource Book (Academic Year 2023-24)*, 2023-05-20 *Learning Elementary Chemistry Class 7 Teacher Resource Book Academic Year 2023 24* *Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science*, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **ATI TEAS Crash Course® Book + Online** John Allen, 2017-06-27 In August 2016 the ATI TEAS exam replaced the TEAS V The ATI TEAS already the most widely used nursing school admission exam in the U S is now also used for allied health program admission nationwide REA s ATI TEAS Crash Course is aligned with the most recent exam changes to the TEAS or Test of Essential Academic Skills Our ATI TEAS Crash Course features easy to read review chapters that cover every exam objective in reading mathematics science and English and language usage The new ATI TEAS allows for one unified test to cover BSN ADN PN and Allied Health programs The ATI TEAS test blueprint has been revised to align with evolving developments in the nursing field and the curriculum creating different points of emphasis on the test In addition the number of questions allotted to each content area has been changed and examinees now may use an on screen calculator *McGraw-Hill's SAT Subject Test: Chemistry, 2ed* Thomas Evangelist, 2009-02-01 We want to help you score high on the SAT Chemistry test We ve put all of our proven expertise into McGraw Hill s SAT Subject Test Chemistry to make sure you re fully prepared for this difficult exam With this book you ll get essential skill building techniques and strategies created by a leading high school chemistry teacher You ll also get 5 full length practice tests hundreds of sample questions and all the facts about the current exam With McGraw Hill s SAT Subject Test Chemistry we ll guide you step by step through your preparation program and give you the tools you need to succeed 5 full length sample tests with complete explanations for every question 40 top test items to remember on exam day A step by step review of all topics covered on the exam Teacher recommended tips and strategies to help you raise your score **McGraw-Hill's SAT Subject Test Chemistry, 3rd Edition** Thomas Evangelist, 2011-02-03 Expert guidance on the Chemistry exam Many colleges and universities require you to take one or more SAT II Subject Tests to demonstrate your mastery of specific high school subjects McGraw Hill s SAT Subject Test Chemistry is written by experts in the field and gives you the guidance you need perform at your best This book includes 4 full length sample tests updated for the latest test format 40 top tips to remember on test day Glossary of tested chemistry terms and formulas Tips and strategies from one of the most popular teachers at the renowned Brooklyn Technical High School Diagnostic test to pinpoint strengths and weaknesses Step by step review of all topics covered on the exam In depth coverage of the lab experiment questions that are

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Table of Contents Chapter 4 Atomic Structure Workbook Answers

1. Understanding the eBook Chapter 4 Atomic Structure Workbook Answers
 - The Rise of Digital Reading Chapter 4 Atomic Structure Workbook Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Atomic Structure Workbook Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 4 Atomic Structure Workbook Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Atomic Structure Workbook Answers
 - Personalized Recommendations
 - Chapter 4 Atomic Structure Workbook Answers User Reviews and Ratings
 - Chapter 4 Atomic Structure Workbook Answers and Bestseller Lists
5. Accessing Chapter 4 Atomic Structure Workbook Answers Free and Paid eBooks
 - Chapter 4 Atomic Structure Workbook Answers Public Domain eBooks
 - Chapter 4 Atomic Structure Workbook Answers eBook Subscription Services
 - Chapter 4 Atomic Structure Workbook Answers Budget-Friendly Options
6. Navigating Chapter 4 Atomic Structure Workbook Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 4 Atomic Structure Workbook Answers Compatibility with Devices
 - Chapter 4 Atomic Structure Workbook Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 4 Atomic Structure Workbook Answers
 - Highlighting and Note-Taking Chapter 4 Atomic Structure Workbook Answers
 - Interactive Elements Chapter 4 Atomic Structure Workbook Answers
8. Staying Engaged with Chapter 4 Atomic Structure Workbook Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 4 Atomic Structure Workbook Answers
9. Balancing eBooks and Physical Books Chapter 4 Atomic Structure Workbook Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 4 Atomic Structure Workbook Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 4 Atomic Structure Workbook Answers
- Setting Reading Goals Chapter 4 Atomic Structure Workbook Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 4 Atomic Structure Workbook Answers
- Fact-Checking eBook Content of Chapter 4 Atomic Structure Workbook Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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