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STUDY OF GAS LAWS

SYLLABUS

- (i) The behaviour of gases under changes of temperature and pressure; explanation in terms of molecular motion (particles, atoms, molecules); Boyle's Law and Charles' Law; absolute zero; gas equation; simple relevant calculations.
- The behaviour of gases under changes of temperature and pressure, explanation in term if molecular motion (particles, atoms, molecules)
 - Boyle's Law: statement, mathematical form, simple calculations.
 - Charles' Law: statement, mathematical form, simple calculations.
 - Absolute zero Kelvin scale of temperature.
 - Gas equation $P_1 V_1 / T_1 = P_2 V_2 / T_2$; simple relevant calculations based on gas equation.
- (ii) Relationship between Kelvin scale and Celsius Scale of temperature; Standard temperature and pressure. Conversion of temperature from Celsius Scale to Kelvin scale and vice versa. Standard temperature and pressure. (Simple calculations)

7.1 INTRODUCTION

The state of matter in which inter-particle attraction is weak and inter-particle space is so large that the particles become completely free to move randomly in the entire available space is known as **Gas**. Gas occupies the entire space of the vessel in which it is kept and so takes the shape of the vessel. All gases show uniform behaviour under similar conditions of temperature and pressure irrespective of their chemical nature or colour or odour.

7.2 BEHAVIOUR AND CHARACTERISTIC PROPERTIES OF GASES

Gases are composed of molecules (particles) that are in constant random motion. Kinetic theory helps in explaining the simple relationship that exists between the pressure, the volume, and the temperature of a gas.

The kinetic molecular theory, explains the behaviour and characteristic properties of gases:

(1) Composition of gases : Gases are made up of tiny particles (molecules) moving in all possible directions at all possible speeds. The molecules are negligibly small in size as compared to the volume occupied by the gas.

(2) Gases have neither a fixed volume nor a fixed shape : There is negligible force of attraction between the particles (gas molecules). Therefore, the particles (gas molecules) are free to move in the entire space available to them, their movement is

restricted only by the walls of the container. Thus, they attain the shape of the containing vessel.

(3) Gases exert pressure in all directions: The moving particles (molecules) of a gas collide with each other and also with the walls of the container. Due to these collisions, gas molecules exert pressure. It has been found, that at a given temperature, time and area, the same number of molecules of a gas strike against the walls of the container. Thus, gases exert the same pressure in all directions.

(4) Gases are highly compressible : There are large inter-particle (inter-molecular) spaces between gas molecules, and this accounts for the high compressibility of gases. On applying pressure, the molecules come closer, thus decreasing the volume of the gas.

(5) Gases are highly expansible : Gases increase in volume on decrease in pressure and increase in temperature.

When pressure on an enclosed gas is reduced, its particles (molecules) move apart, thus increasing their inter-molecular spaces. As a result, the volume of the gas increases.

When an enclosed gas is heated, kinetic energy of its molecules increases. Thus, the molecules start moving faster and farther apart from each other, resulting in an increase in the volume of the gas.

(6) Gases have low density : The number of molecules per unit volume in a gas is very small as

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